

UNCLASSIFIED

AD 464159

DEFENSE DOCUMENTATION CENTER

FOR

SCIENTIFIC AND TECHNICAL INFORMATION

CAMERON STATION ALEXANDRIA, VIRGINIA



UNCLASSIFIED

NOTICE: When government or other drawings, specifications or other data are used for any purpose other than in connection with a definitely related government procurement operation, the U. S. Government thereby incurs no responsibility, nor any obligation whatsoever; and the fact that the Government may have formulated, furnished, or in any way supplied the said drawings, specifications, or other data is not to be regarded by implication or otherwise as in any manner licensing the holder or any other person or corporation, or conveying any rights or permission to manufacture, use or sell any patented invention that may in any way be related thereto.

AFML-TR-65-90
Part I

HIGH STRENGTH - HIGH MODULUS GLASS FIBERS

Robert M. McMarlin, Ralph L. Tiede, F. Munro Veazie
OWENS-CORNING FIBERGLAS CORPORATION

TECHNICAL REPORT AFML-TR-65-90

March 1965

AVAILABLE COPY WILL NOT PERMIT
FULLY LEGIBLE REPRODUCTION.
REPRODUCTION WILL BE MADE IF
REQUESTED BY USERS OF LDC.

Air Force Materials Laboratory
Research and Technology Division
Air Force Systems Command
Wright-Patterson Air Force Base, Ohio

464159

CATALOGED BY: DDC

AS AD 17

464159

NOTICES

When Government drawings, specifications, or other data are used for any purpose other than in connection with a definitely related Government procurement operation, the United States Government thereby incurs no responsibility nor any obligation whatsoever; and the fact that the Government may have formulated, furnished, or in any way supplied the said drawings, specifications, or other data, is not to be regarded by implication or otherwise as in any manner licensing the holder or any other person or corporation, or conveying any rights or permission to manufacture, use, or sell any patented invention that may in any way be related thereto.

Qualified requesters may obtain copies of this report from the Defense Documentation Center (DDC), (formerly ASTIA), Cameron Station, Bldg. 5, 5010 Duke Street, Alexandria 4, Virginia.

DDC release to CFSTI not authorized.

Dissemination outside the Department of Defense or to recipients other than Government Contractors is prohibited.

Copies of this report should not be returned to the Research and Technology Division unless return is required by security considerations, contractual obligations, or notice on a specific document.

FOREWORD

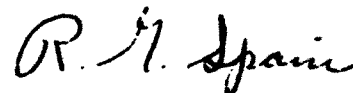
This report was prepared by the Glass Research Laboratory of Owens-Corning Fiberglass Corporation, Granville, Ohio under USAF Contract AF 33(615)-1370. This contract was initiated under Project No. 7320, "Fibrous Materials for Decelerators and Structures", Task No. 732001, "Nonmetallic and Composite Materials" and Project No. 7340, "Organic and Inorganic Fibers", Task No. 734003, "Structural Plastic and Composite Materials". This report covers the work accomplished during the period March 1964 through February 1965.

The contract was initiated by the Nonmetallic Materials Division of the Air Force Materials Laboratory with Walter H. Gloor(MANF) of the Fibrous Materials Branch as Project Engineer.

The work was performed under the technical direction of Mr. F. Munro Veazie, as Program Manager, with Mr. Robert McMarlin and Mr. Ralph Tiede as Project Managers. They were assisted by Messrs. Austin Walpole, Glen Lucas, and John Brookbank.

Manuscript released by authors for publication as an RTD
Technical Report

This technical documentary report has been reviewed and is approved.



R. G. SPAIN, Chief
Fibrous Materials Branch
Nonmetallic Materials Division

ABSTRACT

The results of the present investigation have shown that glass fibers drawn from compositions located in the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-ZnO}$ system possess tensile strengths in excess of 800,000 psi. Glass fibers drawn from compositions in the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-MgO}$ system have given similar results, particularly at fiber diameters in the 0.00015 to 0.00020 inch range.

The modulus of elasticity of glass fibers obtained from compositions located in the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-MgO}$ system can be increased to 16.0×10^6 psi. This is accomplished by lowering the fiber diameter to the 0.00015 to 0.00020 inch range.

In the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-Na}_2\text{O}$ system, the fiber forming temperature, the tensile strength and the modulus of elasticity change with composition in a manner which gives a break in the composition vs. property curves at the point where the Al to Na ratio is 1. This is related to the change in the coordination number of the aluminum at this ratio which has been proposed in the literature.

The properties measured for glasses in the $\text{SiO}_2\text{-TiO}_2\text{-K}_2\text{O}$ system do not appear to lead to as clear a concept of glass structure. Properties vary gradually with composition, with no evidence of a sharp break at any point. There is some indication that the rate of change of properties is dominated more by the silica content than is the case in the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-Na}_2\text{O}$ system.

TABLE OF CONTENTS

	<u>Page</u>
I. Introduction	1
II. Discussion	2
A. Study of $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-R}_x\text{O}$ Glasses.	2
1. Glass Preparation, Fiberization, and Testing . . .	2
2. Tensile Strength and Modulus of Elasticity Results.	2
a. $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-MgO}$ Compositions.	2
b. $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-R}_x\text{O}$ Compositions	14
c. $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-MgO-R}_x\text{O}$ Compositions	15
d. $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-MgO}$ Glass with Additions of other Glass Formers and Intermediates	17
3. The Effect of Fiber Diameter on Tensile Strength and Modulus of Elasticity.	19
B. Relationship Between Composition and Properties of Selected Glasses	19
1. General.	20
2. Experimental Procedure	20
a. $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-Na}_2\text{O}$ System	20
b. $\text{SiO}_2\text{-TiO}_2\text{-K}_2\text{O}$ System	23
3. Experimental Results	27
a. $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-Na}_2\text{O}$ System	27
b. $\text{SiO}_2\text{-TiO}_2\text{-K}_2\text{O}$ System	31
4. Discussion of Results	35
III. Conclusions	44
IV. References	45

LIST OF ILLUSTRATIONS

<u>Figure</u>		<u>Page</u>
1	SiO ₂ -Al ₂ O ₃ -MgO Phase Diagram	3
2	Tensile Strength, Forming Temperature, and Modulus as Effected by the Mol Ratio of Al ₂ O ₃ to MgO	12
3	Tensile Strength, Forming Temperature, and Modulus as Effected by the SiO ₂ Content	13
4	Variation in Tensile Strength and Modulus with R _x O Substitutions for MgO	16
5	Variation in Tensile Strength and Modulus with R _x O Substitutions for 1/2 the MgO on a Mol Basis	18
6	Tensile Strength Versus Al/Na Ratio	24
7	Forming Temperature Versus Al/Na Ratio	25
8	Modulus Versus Al/Na Ratio	26
9	Modulus Versus (Si + Al/O)Ratio.	28
10	Phase Diagram of the SiO ₂ -Al ₂ O ₃ -Na ₂ O System	30
11	Tensile Strength Versus TiO ₂ /K ₂ O Ratio	32
12	Tensile Strength Versus (SiO ₂ + TiO ₂)/K ₂ O Ratio	33
13	Forming Temperature Versus TiO ₂ /K ₂ O Ratio.	34
14	Forming Temperature Versus (SiO ₂ + TiO ₂)/K ₂ O Ratio	36
15	Modulus Versus TiO ₂ /K ₂ O Ratio	37
16	Modulus Versus (SiO ₂ + TiO ₂)/K ₂ O Ratio	38
17	Modulus Versus Per Cent K ₂ O	41
18	Phase Diagram of the SiO ₂ -TiO ₂ -K ₂ O System	41

LIST OF TABLES

<u>Table</u>		<u>Page</u>
1	Theoretical Glass Compositions and Fiber Properties	4 - 10
2	Compositions in Mol Per Cent - $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-Na}_2\text{O}$ System	21
3	Compositions in Mol Per Cent - $\text{SiO}_2\text{-TiO}_2\text{-K}_2\text{O}$ System	23
4	Modulus of As Drawn and Annealed Fibers, $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-Na}_2\text{O}$ System	29

HIGH STRENGTH - HIGH MODULUS GLASS FIBERS

INTRODUCTION

The present study was undertaken for the purpose of preparing and studying glass compositions which would produce glass fibers possessing high tensile strengths and high moduli of elasticity. An additional study was initiated which sought to establish a relationship between tensile strength and the composition and probable constitution of selected glasses. Results of this additional study are presented separately under the title "Relationship Between Composition and Properties of Selected Glasses."

Emphasis in the study of high strength - high modulus glass fibers was placed on glasses of two general classifications. In the first classification, $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-R}_x\text{O}$ glasses were prepared in which the alumina to R_xO ratio was maintained at one on a mol basis. In the second classification, $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-R}_x\text{O}$ glasses were prepared in which the alumina to R_xO ratio varied from 0.5 to 2.0. The R_xO component included the alkali and alkaline earth oxides. Other modifications of both classifications were made using several other glass formers and modifiers.

The study initially proceeded with the preparation of glass compositions in the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-MgO}$ system, notably S-994 and related glasses. Subsequent studies were then made by substituting alkali and alkaline earth oxides for all or part of the MgO in the base $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-MgO}$ glasses.

DISCUSSION

A. Study of $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-R}_x\text{O}$ Glasses1. Glass Preparation, Fiberization, and Testing

Each glass composition was prepared by dry mixing the appropriate end member materials which were of at least U.S.P. grade. The dry batch was melted in precious metal crucibles. Temperatures used were 2900°F for two hours and 3200°F for an additional hour or 2900°F for 24 hours. Both methods gave equivalent results. Each batch was removed from its crucible and remixed at least once during the melting operation. All glasses were cooled by plunging the hot crucible directly from the furnace into a cold water bath. Fibers were obtained from each glass composition by remelting the prepared cullet in a one-hole precious metal bushing capable of reaching temperatures in excess of 3200°F .

To test the fibers, virgin filaments were captured by hand in the usual manner and, to determine tensile strength, were wound on a stainless steel fork. Each fork normally contained 6-8 fiber lengths for study after this operation. The apparatus for measuring tensile strength was designed to test one fiber at a time, recording essentially the load elongation diagram for each fiber. The fibers were loaded by a screw-driven cross-head at a constant rate of strain of 0.25 in./in./minute applied to a 2 inch gage length of fiber. The load cell consisted of a catilever beam with a linear variable differential transformer (LVDT) as the displacement sensing element. The load cell was calibrated by dead weight loading. The LVDT was powered by a strain gage amplifier. The signals from the LVDT were first rectified through matched pairs of germanium diodes, then passed through gain and zero potentiometers and recorded on a single channel recorder (Esterline-Angus).

The fiber diameter of each fiber length (6-8 measurements) was determined from a short length of fiber remaining on the fork after mounting. The fibers were mounted on a microscope slide in a liquid of appropriate index of refraction and their diameters measured under a 430 power microscope with the aid of a micrometer eyepiece.

The modulus of elasticity of a particular glass composition was determined by direct measurement of the elongation of an unannealed fiber under load. In most cases, these fibers had diameters of approximately 0.00060 to 0.00080.

2. Tensile Strength and Modulus of Elasticity Results

- a. Glasses located in the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-MgO}$ system were studied extensively. The composition locations are shown in Figure 1, which represents the phase diagram of this system and shows the boundaries between primary phase fields as straight lines. The compositions studied are located at the circled points. The theoretical glass compositions studied in the system are shown

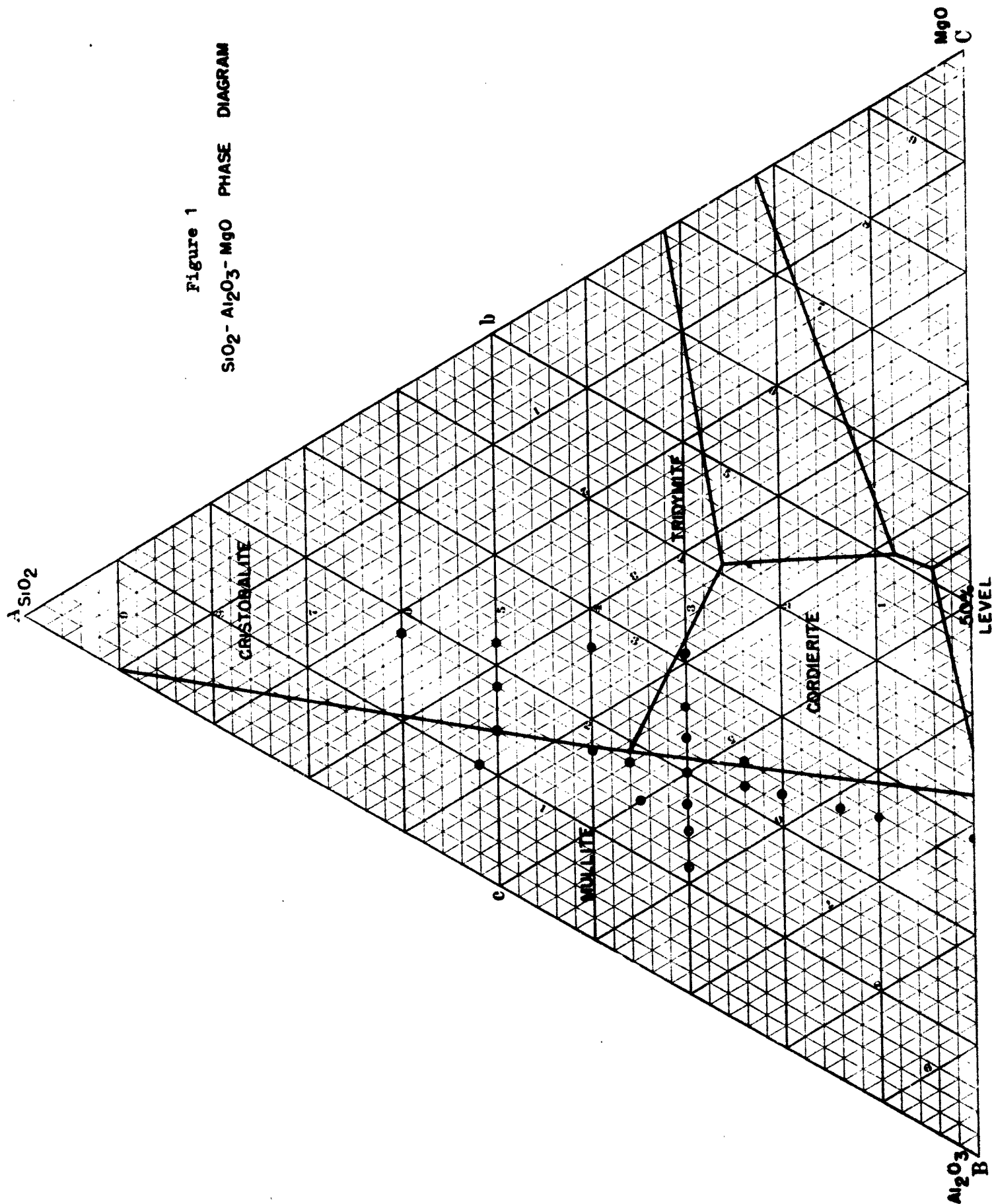


Figure 1
 SiO_2 - Al_2O_3 - MgO PHASE DIAGRAM

TABLE I
THEORETICAL GLASS COMPOSITIONS AND FIBER PROPERTIES

GLASS NO.	OXIDES (a. weight % - b. mol %)										AVERAGE TENSILE STRENGTH x 10 ³ psi	MODULUS x 10 ⁶ psi	AVERAGE DIAMETER x 10 ⁻⁵ in.	FORMING TEMP. °F	NO. OF FIBERS
	SiO ₂	Al ₂ O ₃	MgO	Li ₂ O	K ₂ O	BeO	CaO	ZnO	SrO	BaO					
X-994	a.65.0 b.68.68	25.0 15.56	10.0 15.76								700-600	12.5	16-35	2750	16
X-1751	a.65.0 b.67.7	23.5 14.4	11.5 17.8								674	11.6	31	2700	21
X-1753	a.65.0 b.67.0	22.0 13.3	13.0 19.9								671	12.4	42	2700	20
X-1755	a.65.0 b.69.8	26.4 16.6	8.6 13.7								716	12.0	40	2800	20
X-1757	a.65.0 b.70.5	27.7 17.7	7.3 11.8								678	12.7	35	2850	21
X-1758	a.68.0 b.71.6	23.0 14.2	9.0 14.2								692	12.5	40	2790	21
X-1759	a.70.0 b.73.5	21.5 13.3	8.5 13.3								699	12.6	37	2800	19
X-1760	a.62.0 b.66.0	27.2 17.0	10.8 17.0								834	13.6	32	2700	18
X-1761	a.60.0 b.64.0	28.7 18.0	11.3 18.0								678	12.3	22	27.50	15
X-1762	a.65.0 b.63.9	19.6 12.1	15.4 24.1								600	11.6	37	2600	21
X-1763	a.65.0 b.71.6	29.2 18.9	5.8 9.5								698	12.7	42	2910	19
X-1888	a.55.0 b.59.1	32.3 20.5	12.7 20.4								395	12.3	35	2735	17

TABLE I (Continued)

CLASS NO.	OXIDES (a. weight % - b. mol %)										AVERAGE TENSILE STRENGTH x 10 ³ psi	MODULUS x 10 ⁶ psi	AVERAGE DIAMETER x 10 ⁻⁵ in.	FORMING TEMP. °F.	NO. OF FIBERS
	SiO ₂	Al ₂ O ₃	MgO	Li ₂ O	K ₂ O	BeO	CaO	ZnO	SrO	BaO					
X-1889	a.57.0 b.61.0	30.8 19.4	12.2 19.5								431	12.0	35	2735	16
X-1890	a.50.0 b.54.2	35.9 22.9	14.2 22.9								COULD NOT BE FIBERIZED				
X-1898	a.70.0 b.70.3	16.8 9.9	13.2 19.8								676	11.7	30	2950	6
X-1899	a.75.0 b.75.3	14.0 8.3	11.0 16.5								619	11.1	33	2850	18
X-1900	a.80.0 b.80.2	11.2 6.6	8.8 13.2								651	11.3	33	3145	14
X-1901	a.76.0 b.80.3	19.0 11.8	5.0 7.9								505	-	28	3155	13
X-1902	a.67.5 b.71.8	25.0 15.7	7.5 12.5								726	11.7	16	2915	18
X-1903	a.62.0 b.65.1	26.0 16.1	12.0 18.8								519	12.4	33	2700	17
X-1657	a.75.0 b.78.1	18.0 11.0	7.0 10.9								662	11.5	31	3060	15
X-1659	a.75.0 b.76.6	16.0 9.6	9.0 13.7								716	11.7	32	3170	12
X-1764	a.65.0 b.67.1	27.1 16.5	7.9 16.4								601	10.0	44	2720	21
X-1765	a.65.0 b.64.4	24.3 14.2	10.7 21.4								COULD NOT BE FIBERIZED				
X-1767	a.65.0 b.69.4	29.3 18.4	5.7 12.2								615	10.8	39	2800	21

TABLE I (Continued)

GLASS NO.	OXIDES (a. weight % - b. mol %)										AVERAGE TENSILE STRENGTH	MODULUS x 10 ⁶ psi	AVERAGE DIAMETER x 10 ⁻⁵ in.	FORMING TEMP. °F.	NO. OF FIBERS
	SiO ₂	Al ₂ O ₃	MgO	Li ₂ O	K ₂ O	BaO	CaO	ZnO	SrO	BaO					
X-1768	a. 65.0 b. 70.7	30.6 17.6		4.4 9.7							641	11.0	34	2900	20
X-1769	a. 58.0 b. 70.0	24.8 15.0		7.2 15.0							600	10.0	37	2760	20
X-1770	a. 70.0 b. 71.9	23.2 14.0		6.6 14.1							COULD NOT BE FIBERIZED				
X-1785	a. 62.0 b. 64.2	29.4 17.9		6.6 17.9							546	10.4	33	2710	21
X-1786	a. 60.0 b. 62.2	31.0 18.9		9.0 18.8							337	9.9	49	2640	21
X-1795	a. 65.0 b. 75.2	16.2 12.4		16.6 12.4							COULD NOT BE FIBERIZED				
X-1796	a. 65.0 b. 74.9	12.3 8.4		22.7 16.7							347	7.9	36	2850	19
X-1797	a. 70.0 b. 79.2	15.6 10.4		14.4 10.4							COULD NOT BE FIBERIZED				
X-1798	a. 80.0 b. 86.6	10.4 6.6		9.6 6.8							COULD NOT BE FIBERIZED				
X-1799	a. 65.0 b. 71.0	22.6 14.6		12.4 14.5							486	12.0	24	2750	12
X-1800	a. 65.0 b. 68.8	16.7 10.4		18.3 20.7							456	12.5	49	2660	19
X-1820	a. 65.0 b. 72.8	27.4 18.1		7.6 9.1							583	11.3	38	2940	21
X-1821	a. 70.0 b. 75.4	19.4 12.3		10.7 12.4							598	10.2	35	2810	20

TABLE I (Continued)

GLASS NO.	OXIDES (a. weight % - b. mol %)										AVERAGE TENSILE STRENGTH $\times 10^3$ psi	MODULUS DIAMETER $\times 10^6$ psi $\times 10^{-5}$ in.	FORMING TEMP °F.	NO. OF FIBERS
	SiO ₂	Al ₂ O ₃	MgO	Li ₂ O	K ₂ O	BeO	CaO	ZnO	SrO	BaO				
X-1822	a. 80.0 b. 84.0	12.9 8.0					7.1 8.0							
X-1823	a. 60.0 b. 66.4	25.8 16.8					14.2 16.8				571	11.4	42	19
X-1792	a. 65.0 b. 74.0	19.5 13.1						15.5 13.0			626	11.3	41	17
X-1792 (1)	a. 56.9 b. 68.6	22.9 15.7						18.2 15.7			630	11.8	27	17
X-1792 (2)	a. 59.0 b. 68.68	22.7 15.56						18.3 15.76			815	11.5	26	11
X-1793	a. 75.0 b. 82.1	13.9 8.9						11.1 8.9						
X-1824	a. 65.0 b. 76.0	17.4 12.0							17.6 12.0		432	11.4	40	18
X-1825	a. 65.0 b. 76.0	11.6 8.0							23.5 16.0		355	13.7	29	18
X-1826	a. 65.0 b. 76.0	23.2 16.0							11.8 8.0					
X-1827	a. 60.0 b. 71.9	19.8 14.0							20.2 14.1		568	11.3	37	20
X-1828	a. 70.0 b. 80.0	14.9 10.0							15.1 10.0					
X-1829	a. 80.0 b. 87.3	9.9 6.4							10.1 6.4					
X-1830	a. 65.0 b. 79.8	14.0 10.1												
									21.0 10.1	393		45	3200	7

TABLE I (Continued)

GLASS NO.	OXIDES (a. weight % - b. mol %)										AVERAGE TENSILE STRENGTH $\times 10^3$ psi	AVERAGE DIAMETER $\times 10^{-5}$ in.	FORMING TEMP. °F.	NO. OF FIBERS
	SiO ₂	Al ₂ O ₃	MgO	Li ₂ O	Na ₂ O	K ₂ O	BeO	CaO	ZnO	SrO	BaO			
X-1831	a. 65.0 b. 80.7	8.8 6.4									26.3 12.8	226 8.7	31 3115	16
X-1851	a. 65.0 b. 66.3	28.1 16.8					6.9 16.9					682 13.7	57 3000	16
X-1852	a. 65.0 b. 61.1	23.5 13.0					11.5 26.0					327 14.0	34 3050	12
X-1891	a. 70.0 b. 71.2	24.1 14.4					5.9 14.4					635 13.2	39 3025	17
X-1835	a. 65.8 b. 68.7	24.4 15.6	5.1 7.9	3.7 7.9								407 12.1	36 2735	17
X-1836	a. 63.3 b. 68.7	24.4 15.6	4.9 7.9	7.5 7.9								532 11.0	40 2850	15
X-1837	a. 60.9 b. 68.7	23.5 15.6	4.7 7.9	11.0 7.4								413 10.6	42 2955	17
X-1838	a. 63.7 b. 68.7	24.5 15.6	4.9 7.9				6.8 7.9					604 12.6	40 2795	17
X-1839	a. 61.8 b. 68.7	23.8 15.6	4.8 7.9				9.6 7.9					542 12.9	39 2955	16
X-1840	a. 60.2 b. 68.7	23.2 15.6	4.6 7.9				11.9 7.9					391 11.7	43 2900	18
X-1841	a. 57.0 b. 68.7	21.9 15.6	4.4 7.9				16.7 7.9					485 -	44 2900	16
X-1896	a. 66.2 b. 68.6	25.5 15.6	5.1 7.9				3.2 7.9					638 12.4	33 2850	16
X-1897	a. 71.7 b. 68.6	13.6 7.8	11.0 15.8				3.4 7.8					504 11.7	33 2795	16

TABLE I (Continued)

GLASS NO.	OXIDES (a. weight % - b. mol %)								AVERAGE TENSILE STRENGTH x 10 ³ psi	MODULUS x 10 ⁶ psi	AVERAGE DIAMETER x 10 ⁻⁵ in.	FORMING TEMP. °F.	NO. OF FIBERS
	SiO ₂	Al ₂ O ₃	MgO	P ₂ O ₅	TiO ₂	ZrO ₂	HfO ₂	GeO ₂					
X-1842	a. 65.0 b. 74.8	18.0 11.6	7.0 11.7	10.0 4.9					488	11.6	38	2950	18
X-1843	a. 60.0 b. 68.6	18.0 12.1	7.0 12.0	15.0 7.3					357	11.4	35	2900	16
X-1844	a. 55.0 b. 65.1	18.0 12.5	7.0 12.4	20.0 10.0					COULD NOT BE FIBERIZED				
X-1845	a. 55.0 b. 61.9	25.0 16.6	10.0 16.8	10.0 4.7					580	11.8	29	2900	14
X-1849	a. 64.0 b. 68.1	25.0 15.7	10.0 15.8	1.0 0.4					554	12.0	34	2850	18
X-1850	a. 60.0 b. 65.4	25.0 16.1	10.0 16.3	5.0 2.3					406	11.4	38	2735	17
X-1856	a. 50.0 b. 58.1	25.0 17.3	10.0 17.1	15.0 7.4					494	11.7	46	2900	20
X-1846	a. 60.0 b. 64.2	25.0 15.8	10.0 15.9		5.0 4.1				361	12.0	39	2735	16
X-1847	a. 55.0 b. 59.7	25.0 16.0	10.0 16.2		10.0 8.2				260		32	2700	15
X-1846	a. 63.0 b. 66.9	25.0 15.6	10.0 15.8		2.0 1.6				393	12.3	34	2775	16
X-1853	a. 64.5 b. 68.3	25.0 15.6	10.0 15.8			0.5 0.3			620	11.6	35	2735	20
X-1854	a. 64.0 b. 68.0	25.0 15.6	10.0 15.8			1.0 0.5			626	12.1	35	2735	17
X-1855	a. 64.0 b. 69.1	22.2 14.1	8.8 14.1			5.0 2.7			585	12.5	34	2795	17

TABLE I (Continued)

GLASS NO.	OXIDES (a. weight % - b. mol %)							AVERAGE TENSILE STRENGTH x 10 ³ psi	MODULUS x 10 ⁶ psi	AVERAGE DIAMETER x 10 ⁻⁵ in.	FORMING TEMP. ° F.	NO. OF FIBERS
	SiO ₂	Al ₂ O ₃	MgO	P ₂ O ₅	TiO ₂	ZrO ₂	HfO ₂					
X-1857	a. 62.0	25	10			3		511	13.5	31	2770	17
	b. 66.6	15.8	16.0			1.5						
X-1858	a. 60.0	25	10			5		524	12.1	-	2750	21
	b. 65.1	16.0	16.2			2.7						
X-1859	a. 63.0	19.4	7.6			10.0		615	12.8	39	2975	18
	b. 69.5	12.6	12.5			5.4						
X-1893	a. 62.0	17.9	7.1			13.0		429	11.6	51	3100	20
	b. 69.3	11.8	11.8			7.1						
X-1894	a. 65.0	17.9	7.1			10.0		457	11.7	37	3100	20
	b. 71.4	11.6	11.6			5.3						
X-1887	a. 64	25	10				1	501	12.0	26	2955	17
	b. 68.1	15.7	15.9				0.3					
X-1860	a. 64.0	25.0	10.0				1	477	12.7	43	3095	14
	b. 67.9	15.6	15.8				0.6					
X-1885	a. 60.0	25.0	10.0				5.0	508	12.2	32	2850	21
	b. 64.8	15.9	16.0				3.1					
X-1886	a. 55.0	25.0	10.0				10.0	467	12.6	39	2850	19
	b. 60.8	16.3	16.4				6.4					

in Table 1, which also lists pertinent physical properties for each glass. On Table 1, see glass numbers X-994, X-1751 through X-1763, X-1888 through X-1890, X-1898 through X-1903, X-1657 and X-1659.

Most of the data presented describes glass compositions with either constant weight percentages of SiO_2 and variable Al_2O_3 to MgO ratios or constant Al_2O_3 to MgO ratios and variable weight percentages of SiO_2 . These data are summarized graphically in Figures 2 and 3.

Figure 2 indicates that at a constant SiO_2 content of 65% by weight, the tensile strength of the series of compositions increases with an increasing Al_2O_3 to MgO ratio. The increase continues until a maximum is reached at an Al_2O_3 to MgO ratio of 1. At this point it is postulated that additional Al_2O_3 enters sixfold coordination in the glass structure.⁽¹⁾ The data show a reverse change in slope at this point in the curve, which indicates such a structural change is possible and does take place. Increasingly higher forming temperatures and much greater devitrification tendencies above the Al_2O_3 to MgO ratio of 1.0 also make it difficult to interpret the data obtained, but there is a good indication that beyond 1.5 mol % the tensile strength is beginning to increase again at a lower rate.

Tensile strength results for X-994 as listed in Table 1 show a range of values. It is felt that these results are indicative of the structural sensitivity of the glass at this point ($\text{Al}_2\text{O}_3/\text{MgO} = 1.0$). Strength values for other glasses at this ratio have shown similar variability and reinforce the view that slight changes in composition produce fairly large changes in strength.

The modulus values plotted in Figure 2 show no clear-cut trend. There appears to be an over-all increase in modulus with increasing Al_2O_3 to MgO ratio.

Figure 3 indicates that at a constant Al_2O_3 to MgO ratio of 1.0 and a variable SiO_2 content by weight the tensile strength of the series of compositions increases with an increase in the SiO_2 content. The increase is continuous up to a SiO_2 level of from 60-65% where the tensile strength begins to drop again. Again increased forming temperatures and greater devitrification tendencies are found at the higher SiO_2 levels. The plotted value beyond 70% by weight SiO_2 may therefore not be entirely correct.

The high tensile strength peak for X-1760 and the range of values for X-994 are again indicative of the composition sensitivity of these glasses at this point.

The modulus values appear to follow the trend established by the strength values, with maximum values located in the 60-70% SiO_2 range.

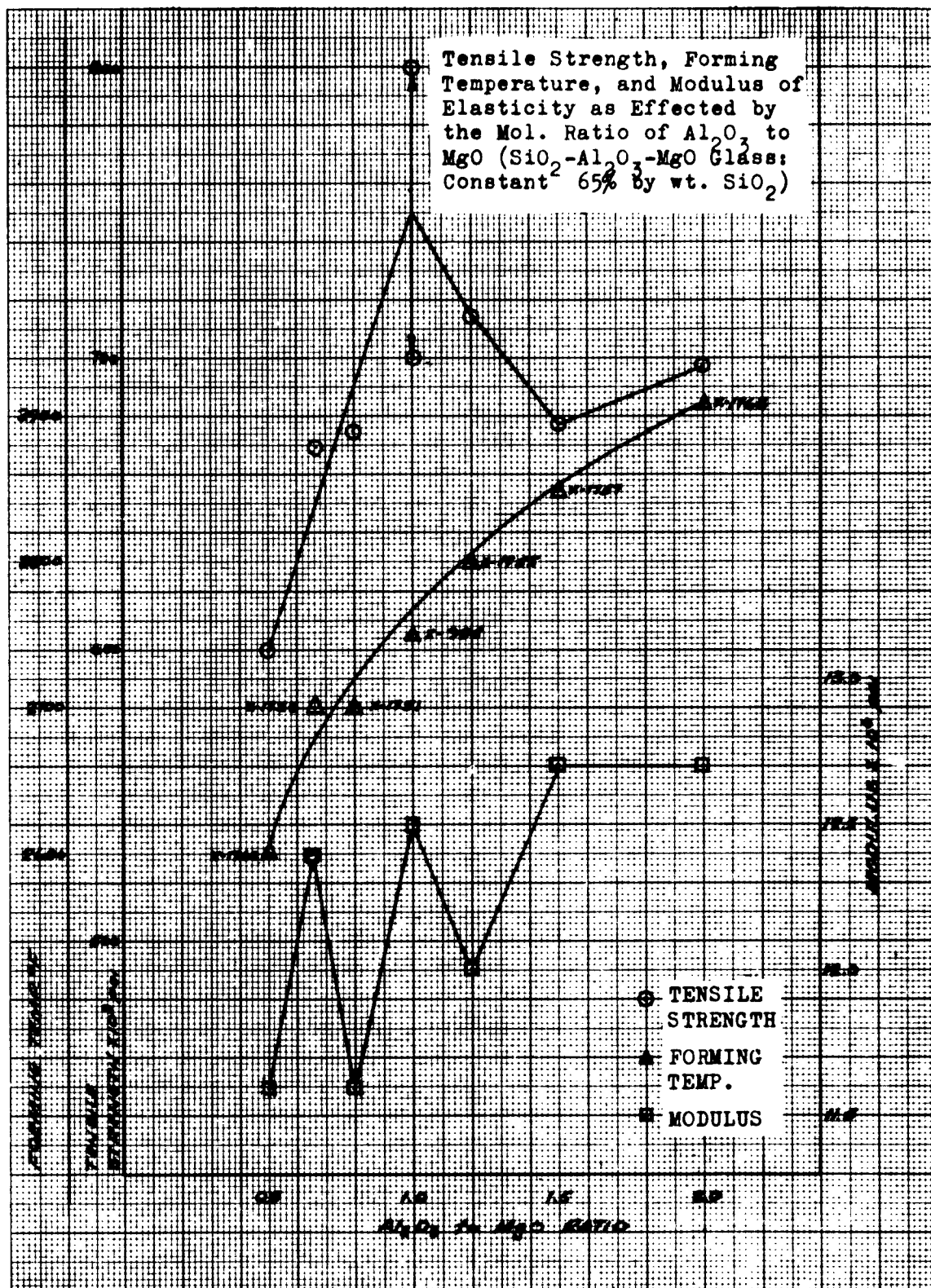


FIGURE 2 Tensile Strength, Forming Temperature, and Modulus as Effected bu the Mol Ratio of Al_2O_3 to MgO

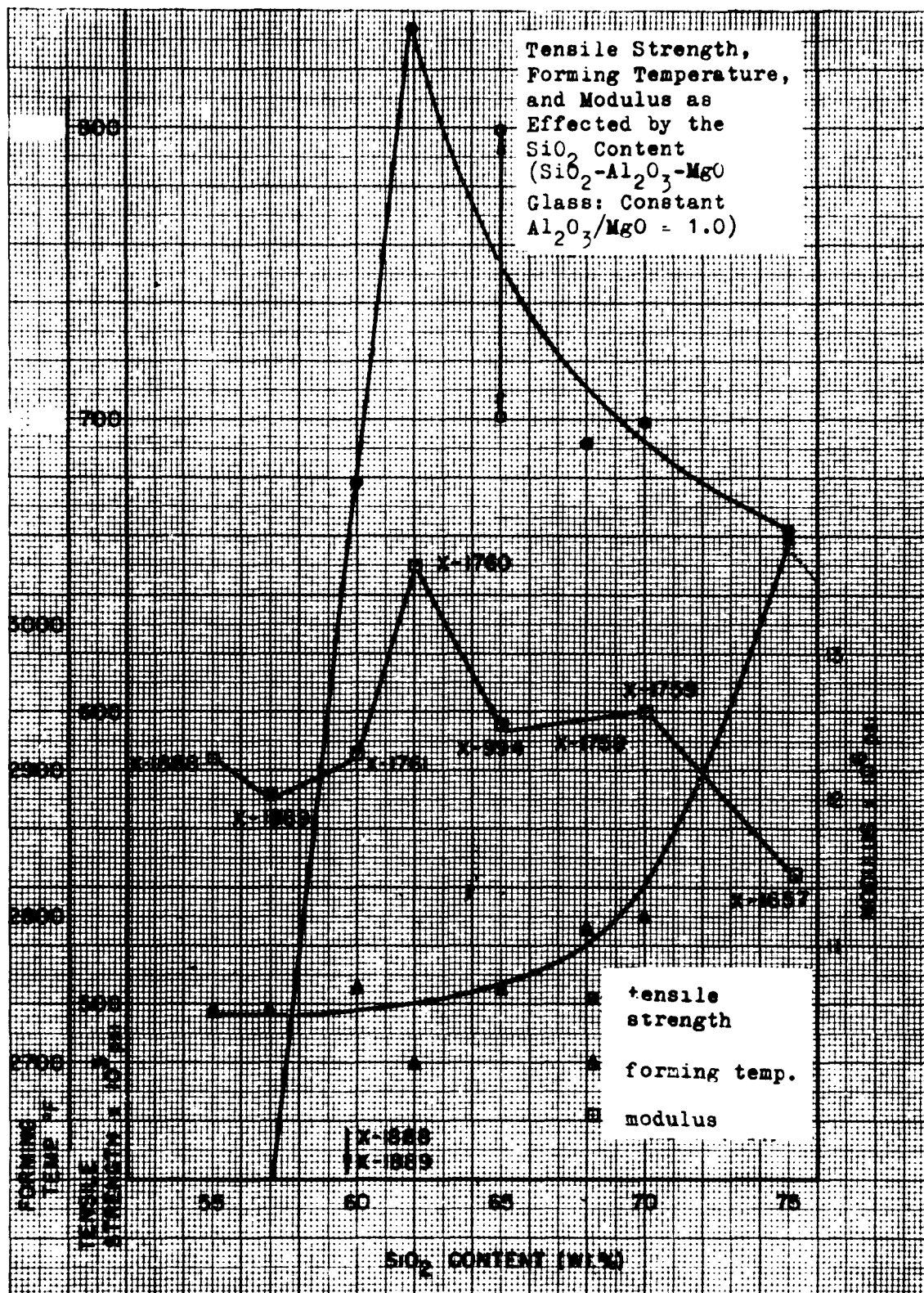


FIGURE 3 Tensile Strength, Forming Temperature, and Modulus as Effected by the SiO_2 Content

The tensile strength values of the remaining glasses which are not plotted tend to point out the trends established in Figures 2 and 3. These are as follows:

- (1) The glasses with the best forming characteristics and highest tensile strengths are located at an approximate Al_2O_3 to MgO ratio of approximately 1.0 and at a SiO_2 level ranging from 60-70% by weight.
- (2) Glasses with Al_2O_3 to MgO ratios much less than 1.0 do not have desirable tensile strengths.
- (3) Glasses with Al_2O_3 to MgO ratios much greater than 1.0 do not have high strengths because of fiberization difficulties and greater devitrification tendencies.
- (4) Glasses within the desirable composition limits (1) have tensile strength values which are sensitive to slight structural changes because of the changes in coordination of the Al^{+++} ion in this area. It is therefore postulated that there is a structural link between the coordination of the Al^{+++} ion in the glass and the tensile strength and modulus of filaments made from the glass.

b. SiO_2 - Al_2O_3 - R_xO Compositions

Table 1 lists the theoretical compositions and properties of a series of glasses containing SiO_2 , Al_2O_3 , and R_xO (where R_xO is Li_2O , K_2O , BeO , CaO , ZnO , SrO , or BaO .) Attention is directed to glasses numbered X-1764 through X-1770, X-1785 and X-1786, X-1795 through X-1796, X-1799 and X-1800, X-1820 through X-1823, X-1824 through X-1831, X-1851 and X-1852, and X-1891 through X-1893.

Figure 4 summarizes the properties of fibers produced from these glasses. The highest average strength and modulus value in each series of compositions has been compared to the range of values established for the SiO_2 - Al_2O_3 - MgO series. The tensile strength values plotted for the alkali oxide substitutions show a continuous strength decreases in going from MgO to Li_2O to K_2O . A similar decreases occurs in the modulus values. It is postulated that the Li^+ ion causes a reduction because its valence is only one and it has a much lower field strength than Mg^{++} . Similarly, the K^+ ion has a valence of only one, has an even lower field strength than Li^+ , and has an ionic radius about twice as large as Mg^{++} .

The alkaline earth oxide substitutions for MgO show a continuous decreases in strength. Ranking the alkaline earth ions by ionic size and field strength may help account for the trend in the tensile strength.

<u>Ion</u>	<u>Ionic Size</u>	<u>Field Strength</u>
Be ⁺⁺	0.31	16.7
Mg ⁺⁺	0.65	4.5
Zn ⁺⁺	0.74	3.6
Ca ⁺⁺	0.99	2.0
Sr ⁺⁺	1.13	1.6
Ba ⁺⁺	1.35	1.1

The Zn⁺⁺ ion, though not an alkaline earth ion, is included to show its similarity to Mg⁺⁺. It would thus be expected to act like Mg⁺⁺. This is the case, as shown in Figure 4, at least for tensile strength values.

With such a trend established, it would be expected that a SiO₂-Al₂O₃-BeO glass would show outstanding strength properties. There is some indication that such may be the case, but no final values are yet available. Glasses of this type have enhanced devitrification and phase separation tendencies, which would be explained by the extremely high field strength of Be⁺⁺. Currently, such tendencies have made fiberization of good quality fibers difficult.

The plot of modulus of elasticity values for the alkaline earth oxide substitutions show a variable pattern. The BeO addition produces the best results, with SrO, surprisingly, a close second. Unfortunately, the high modulus values do not correspond to the strongest glasses in the series.

Summarizing the results of the study in which alkali oxides and alkaline earth oxides were substituted for MgO in the SiO₂-Al₂O₃-MgO series of glasses, it should be noted that:

- (1) The substitution of ZnO for MgO produced the only glasses with tensile strength values comparable to those in the SiO₂-Al₂O₃-MgO system.
- (2) When substituting alkali oxides for MgO, the tensile strength decreased as the size of the ion increased.
- (3) When substituting alkaline earth oxides for MgO, the tensile strength decreased as the size of the ion increased. The effect of BeO has not yet been definitely determined and should not be included in this conclusion.
- (4) None of the oxide substitutions for MgO improved the modulus of elasticity above that associated with glasses in the SiO₂-Al₂O₃-MgO system, except for the BeO substitution.

This is attributed to the small size and high field strength of the Be⁺⁺ ion.

c. SiO₂-Al₂O₃-MgO-R_xO Compositions

Table 1 lists the theoretical compositions and properties of a series of glasses based on the X-994 composition where R_xO

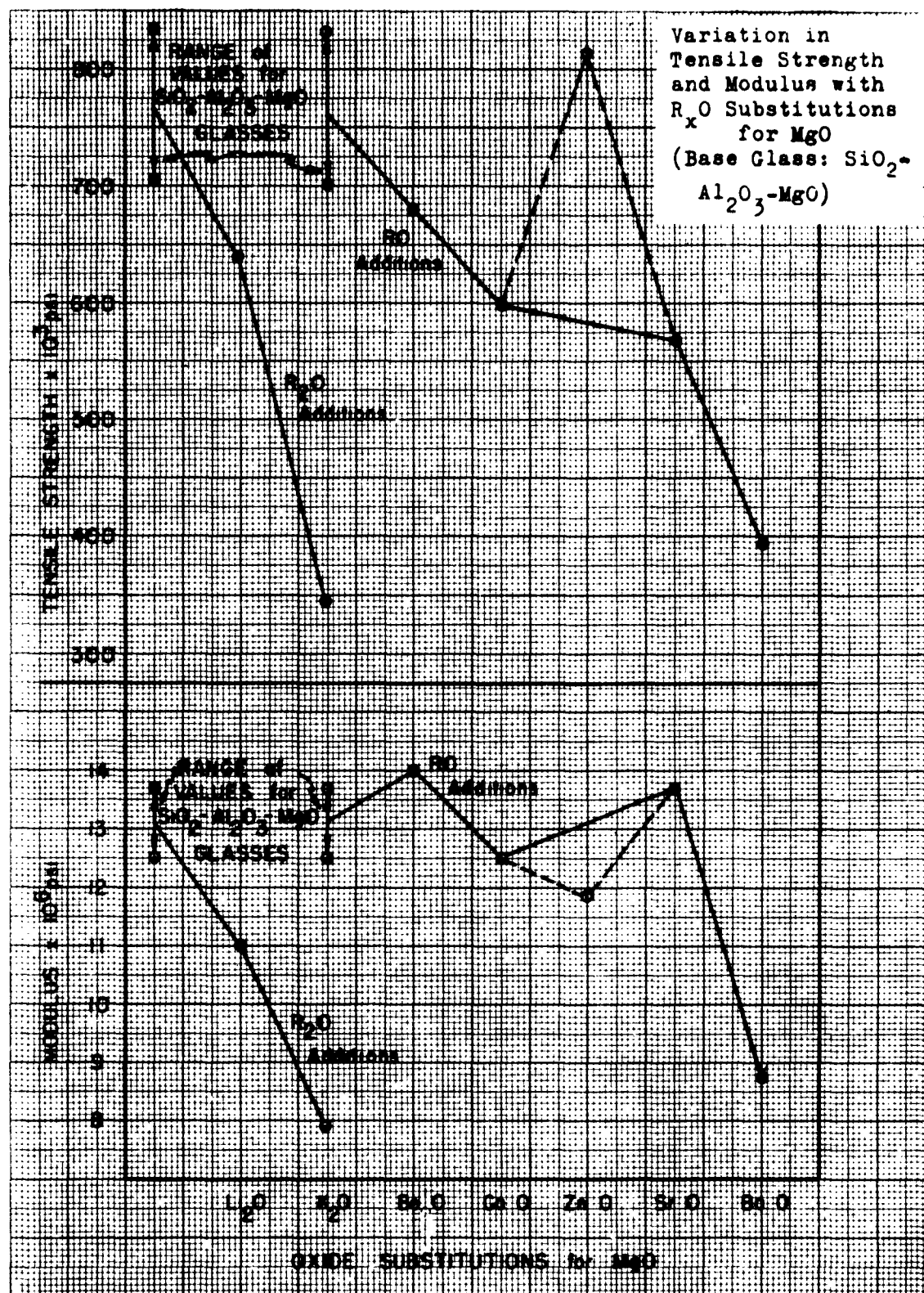


FIGURE 4 Variation in Tensile Strength and Modulus with R_xO Substitutions for MgO

component is substituted for half the MgO on a mol basis.
($R_xO = Li_2O, Na_2O, K_2O, BeO, CaO, ZnO, SrO, BaO.$) Attention
is directed to glasses numbered X-1835 through X-1841 and X-1896.

Figure 5 summarizes the property data. The value of tensile strength and modulus for each oxide addition is compared to the range of values established for the $SiO_2-Al_2O_3-MgO$ glass series. No trend is evident except to note that all of the additions tend to reduce the tensile strength of the $SiO_2-Al_2O_3-MgO$ series of glasses. The trend of the modulus curve cannot be explained, except to note that the glasses with higher modulus values do not correspond to those with high tensile strengths, except in the case of the $SiO_2-Al_2O_3-MgO-BeO$ glass.

1. $SiO_2-Al_2O_3-MgO$ Glasses with Additions of Other Glass Formers and Intermediates

Table 1 lists the theoretical compositions and properties of a series of glasses containing $SiO_2-Al_2O_3-MgO$ and one other glass forming or intermediate oxide ($P_2O_5, TiO_2, ZrO_2, HfO_2$, or GeO_2). On Table 1 see the following glass numbers:

Glasses containing P_2O_5 - X-1842 through X-1845, X-1849 through X-1850, and X-1856.

Glasses containing TiO_2 - X-1846 through X-1848.

Glasses containing ZrO_2 - X-1853 through X-1855, X-1857 through X-1859, X-1893 through X-1894

Glasses containing HfO_2 - X-1867

Glasses containing GeO_2 - X-1860, X-1865, X-1886

The X-994 composition was used as the base glass. In general, the added oxide was either substituted for part of the SiO_2 content of the X-994 composition or for part of the SiO_2, Al_2O_3 , and MgO content on a weight basis. In all cases the mol ratio of Al_2O_3 to MgO remained approximately equal to 1.

A systematic study of the exact effect of each addition on the tensile strength of the base glass was not attempted. In general, however, the tensile strength results indicate that the addition of P_2O_5 for a portion of the SiO_2 resulted in a general strength decrease as the P_2O_5 content increased. An even greater strength drop resulted when P_2O_5 was substituted for a portion of the SiO_2, Al_2O_3 , and MgO . Modulus values were not greatly affected by any of the additions and cannot be explained.

TiO_2 substitutions for part of the SiO_2 caused the greatest drop in strength of all the additions used. The strength decreased as the TiO_2 content increased. Again, modulus values were not greatly affected and this response cannot be explained.

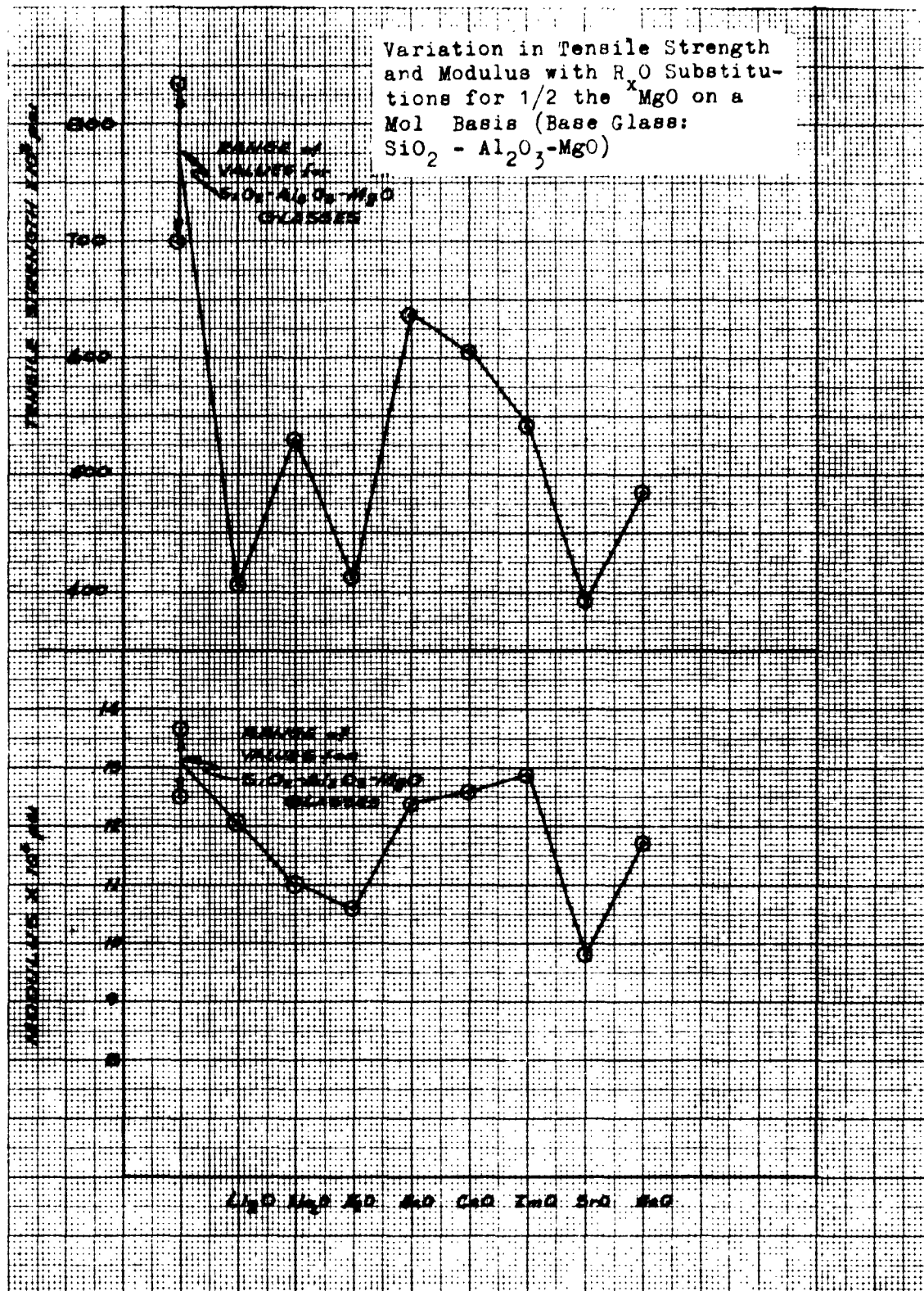


FIGURE 5 Variation in Tensile Strength and Modulus with R_2O Substitutions for $1/2$ the MgO on a Mol Basis

Interpretation of the data for the glasses containing ZrO_2 is difficult. However, it appears that small ZrO_2 additions improve modulus values slightly. Even such small additions, however, decrease tensile strength. HfO_2 has the same effect.

Finally, the addition of GeO_2 for part of the SiO_2 greatly reduces the tensile strength below that associated with X-994. Modulus does not seem to be affected, but again no explanation is offered.

3. The Effect of Fiber Diameter on Tensile Strength and Modulus of Elasticity

Preliminary results from several studies of the forming parameters associated with the fiberization of SiO_2 - Al_2O_3 - MgO glasses have indicated that tensile strength and modulus properties were improved when the diameter of fibers under test was reduced. The tensile strength of virgin X-994 fiber is normally 700-725,000 psi for a 0.00040 inch diameter fiber. This figure is apparently raised to 800-850,000 psi for a fiber with an apparent diameter of 0.00015 to 0.00020 inch. Since there are problems encountered with accurately measuring fibers of this diameter with the micrometer eyepiece, further work in this area is warranted.

Modulus of elasticity values are also apparently improved for lower diameter fibers. The following modulus measurements have been made:

<u>Glass</u>	<u>Modulus $\times 10^6$ psi</u>	<u>Method Employed</u>	<u>Measured Diameter</u>
X-994	12.4	Elongation	0.00065 in.
X-994	13.4	Elongation	0.00031
X-994	13.6	Sonic	0.00027
X-994	16.0	Elongation	0.00018

Note the modulus value, measured by the sonic method, of a fiber with a diameter of 0.00027 inches. The sonic method does not require a diameter measurement. This finding reinforces the view that the lower diameter fibers do have improved moduli. Further work in this area is definitely desirable.

Glasses in systems other than the SiO_2 - Al_2O_3 - MgO series have been checked to determine the effect of reductions in fiber diameter on both tensile strength and modulus. The properties of glasses in the SiO_2 - Al_2O_3 - Li_2O and SiO_2 - Al_2O_3 - CaO systems did not change with reductions in fiber diameter. Improved properties were evident, however, in the SiO_2 - Al_2O_3 - ZnO system.

B. Relationship Between Composition and Properties of Selected Glasses

1. General

Glass fibers with the highest possible tensile strength and modulus of elasticity are desired for many important applications. These properties are intimately related to the ultimate structure of glass. Knowledge of the structure of a glass and its relationship

These properties are invaluable in developing new glass compositions with superior performance.

The work reported here was undertaken to provide some insight into the relationship between glass structure and the properties of tensile strength and modulus of elasticity. In papers by Day and Rindone^(1,2), it has been shown that a number of physical properties of certain glasses in the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-Na}_2\text{O}$ field vary in a systematic manner with composition changes. Further, composition vs. property curves usually showed breaks which were related to proposed changes in glass structure. It was found that in a set of related compositions in which the Al to Na ratio varied, a break occurred at the point where this ratio was one. It was postulated that at ratios below one, the aluminum was in fourfold coordination and participated in the glass network, while at higher ratios some of the aluminum assumed sixfold coordination. This proposition was based in part upon the total number of oxygen ions available. At Al to Na ratios less than one, there are less than two. The work reported here started as an attempt to find out whether modulus of elasticity and tensile strength of fibers could be related to glass composition and presumably structure in the same manner as had been done with other properties.

Considerable property data have also been published for glasses in the $\text{SiO}_2\text{-TiO}_2\text{-K}_2\text{O}$ system⁽³⁾. These glasses were also reported to show deviations in curves of properties vs. composition which were tentatively related to changes in coordination number of the titanium from six to eight. These observations, plus the superficial similarity between the two systems, led to the extension of the work by including a study of this composition field also.

The $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-K}_2\text{O}$ system is an attractive one for further study, inasmuch as it differs from each of the previously studied systems by only one constituent. It may be possible to relate the two through study of this system.

II. Experimental Procedure

a. $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-Na}_2\text{O}$ System

(1) Compositions Studied

Table II gives the compositions in mol percentages for all glasses studied. The glasses are divided into five related series. Series II and IV comprise sets of compositions similar to those given the same designation in the papers by Day and Rindone^(1,2). In general, all glasses in each series contain the same percentage of silica. The exception is Series IV in which all components vary simultaneously.

(2) Preparation and Testing of Specimens

Batches compounded from relatively pure commercial raw materials were melted in crucibles to form the glass used to

TABLE 1

COMPOSITIONS IN MOL PER CENT
SiO₂-Al₂O₃-Na₂O SYSTEMSeries II

	<u>X-1580A</u>	<u>X-1585A</u>	<u>X-1586A</u>	<u>X-1587A</u>	<u>X-1588A</u>	<u>X-1589A</u>	<u>X-1590A</u>
SiO ₂	75.00	75.00	75.00	75.00	75.00	75.00	75.00
Al ₂ O ₃	-	3.13	6.25	9.38	12.50	15.00	16.25
Na ₂ O	25.00	21.88	18.75	15.63	12.50	10.00	8.75
Al/Na	0	.14	.33	.60	1.00	1.50	1.86

Series IV

	<u>X-1580A</u>	<u>X-1591B</u>	<u>X-1592B</u>	<u>X-1593A</u>	<u>X-1594A</u>	<u>X-1595A</u>
SiO ₂	75.00	71.43	66.67	65.52	64.29	62.96
Al ₂ O ₃	-	7.14	16.67	18.97	21.43	24.07
Na ₂ O	25.00	21.43	16.67	15.52	14.29	12.96
Al/Na	0	.33	1.00	1.22	1.50	1.86

Series V

	<u>X-1866</u>	<u>X-1867</u>	<u>X-1868</u>	<u>X-1869</u>	<u>X-1870</u>	<u>X-1871</u>	<u>X-1961</u>
SiO ₂	65.97	65.97	65.97	65.97	65.97	65.97	65.97
Al ₂ O ₃	-	6.00	12.00	17.00	20.00	22.00	24.00
Na ₂ O	34.03	28.03	22.03	17.03	14.03	12.03	10.03
Al/Na	0	.21	.55	1.00	1.43	1.83	2.40

Series VI

	<u>X-1877</u>	<u>X-1878</u>	<u>X-1879</u>	<u>X-1970</u>	<u>X-1971</u>	<u>X-1880</u>	<u>X-1972</u>
SiO ₂	57.00	57.00	57.00	57.00	57.00	57.00	57.00
Al ₂ O ₃	-	6.00	12.00	14.09	15.48	17.00	18.40
Na ₂ O	43.00	37.00	31.00	28.91	27.52	26.00	24.60
Al/Na	0	.16	.39	.49	.56	.65	.75

	<u>X-1881</u>	<u>X-1882</u>	<u>X-1883</u>	<u>X-1962</u>	<u>X-1884</u>
SiO ₂	57.00	57.00	57.00	57.00	57.00
Al ₂ O ₃	21.50	24.50	27.50	29.00	30.50
Na ₂ O	21.50	18.50	15.50	14.00	12.50
Al/Na	1.00	1.33	1.78	2.07	2.44

TABLE II - (Continued)

	<u>Series VII</u>						
	<u>X-1938</u>	<u>X-1939</u>	<u>X-1940</u>	<u>X-1941</u>	<u>X-1942</u>	<u>X-1943</u>	<u>X-1944</u>
SiO ₂	50.00	50.00	50.00	50.00	50.00	50.00	50.00
Al ₂ O ₃	-	6.00	12.00	18.00	21.50	25.00	29.00
Na ₂ O	50.00	44.00	38.00	32.00	28.50	25.00	21.00
Al/Na	0	.14	.32	.56	.76	1.00	1.38

	<u>Series VII - (Continued)</u>	
	<u>X-1958</u>	<u>X-1945</u>
SiO ₂	50.00	50.00
Al ₂ O ₃	31.00	33.00
Na ₂ O	19.00	17.00
Al/Na	1.63	1.94

charge the bushing. Fibers between 0.00035 and 0.00040 inch in diameter were formed for tensile strength measurements. Virgin fibers were collected on forks, placed in a desiccator, and tested promptly in the manner previously described. At least three sets of fibers sufficient to give fifteen breaks each were tested for each glass. Results reported are usually the average of the results for the two closest sets. In a few cases, additional sets of fibers were tested, in which case the averages reported are for more than two sets.

Modulus of elasticity measurements were made on fibers in the 0.00075 to 0.00085 inch diameter range by measuring the elongation under load. This size of fiber was selected because it was convenient to form and to measure. Some difficulty was found in obtaining reproducible results with fibers varying widely in fiber diameter. Although the source of difficulty was never completely isolated, it was found that reproducibility was reasonably good when measurements were performed on fibers of this selected diameter range only. Each point reported is the average of at least eight measurements on four fibers.

- (3) A rough measurement of relative viscosities of the glasses was made by noting the bushing temperature at which fibers were formed. Results are approximate only, since no effort was made to calibrate thermocouples, or to compensate for any drift in thermocouple output that may have occurred as they aged. Also, several different bushings were used, with possible bushing-to-bushing variation.

TABLE III

COMPOSITIONS IN MOL PER CENT
SiO₂-TiO₂-K₂O SYSTEM

	<u>X-1926</u>	<u>X-1927</u>	<u>X-1928</u>	<u>X-1929</u>	<u>X-1930</u>	<u>X-1931</u>	<u>X-1932</u>
SiO ₂	60.0	60.0	60.0	60.0	60.0	60.0	60.0
TiO ₂	-	5.0	10.0	15.0	20.0	25.0	30.0
K ₂ O	40.0	35.0	30.0	25.0	20.0	15.0	10.0
	<u>X-1946</u>	<u>X-1947</u>	<u>X-1948</u>	<u>X-1949</u>	<u>X-1950</u>	<u>X-1951</u>	
SiO ₂	70.0	70.0	70.0	70.0	70.0	70.0	
TiO ₂	-	5.0	10.0	15.0	20.0	25.0	
K ₂ O	30.0	25.0	20.0	15.0	10.0	5.0	
	<u>X-1977</u>	<u>X-1976</u>	<u>X-1979</u>	<u>X-1980</u>			
SiO ₂	80.0	80.0	80.0	80.0			
TiO ₂	-	5.0	10.0	15.0			
K ₂ O	20.0	15.0	10.0	5.0			
	<u>X-1981</u>	<u>X-1982</u>	<u>X-1983</u>	<u>X-1984</u>	<u>X-1985</u>	<u>X-1986</u>	<u>X-1987</u>
SiO ₂	50.0	50.0	50.0	50.0	50.0	50.0	50.0
TiO ₂	40.0	35.0	30.0	25.0	20.0	15.0	10.0
K ₂ O	10.0	15.0	20.0	25.0	30.0	35.0	40.0
	<u>X-1992</u>	<u>X-1993</u>	<u>X-1994</u>	<u>X-1995</u>	<u>X-1996</u>	<u>X-1997</u>	<u>X-1998</u>
SiO ₂	40.0	40.0	40.0	40.0	40.0	40.0	40.0
TiO ₂	10.0	15.0	20.0	25.0	30.0	35.0	40.0
K ₂ O	50.0	45.0	40.0	35.0	30.0	25.0	20.0
	<u>X-1999</u>	<u>X-2000</u>					
SiO ₂	40.0	40.0					
TiO ₂	45.0	50.0					
K ₂ O	15.0	10.0					

b. SiO₂-TiO₂-K₂O System(1) Compositions Studied

Table III gives the compositions studied in mol percentages.

(2) Preparation and Testing of Specimens

Glasses were prepared, fiberized, and tested as described for the glasses in the silica-alumina-soda system.

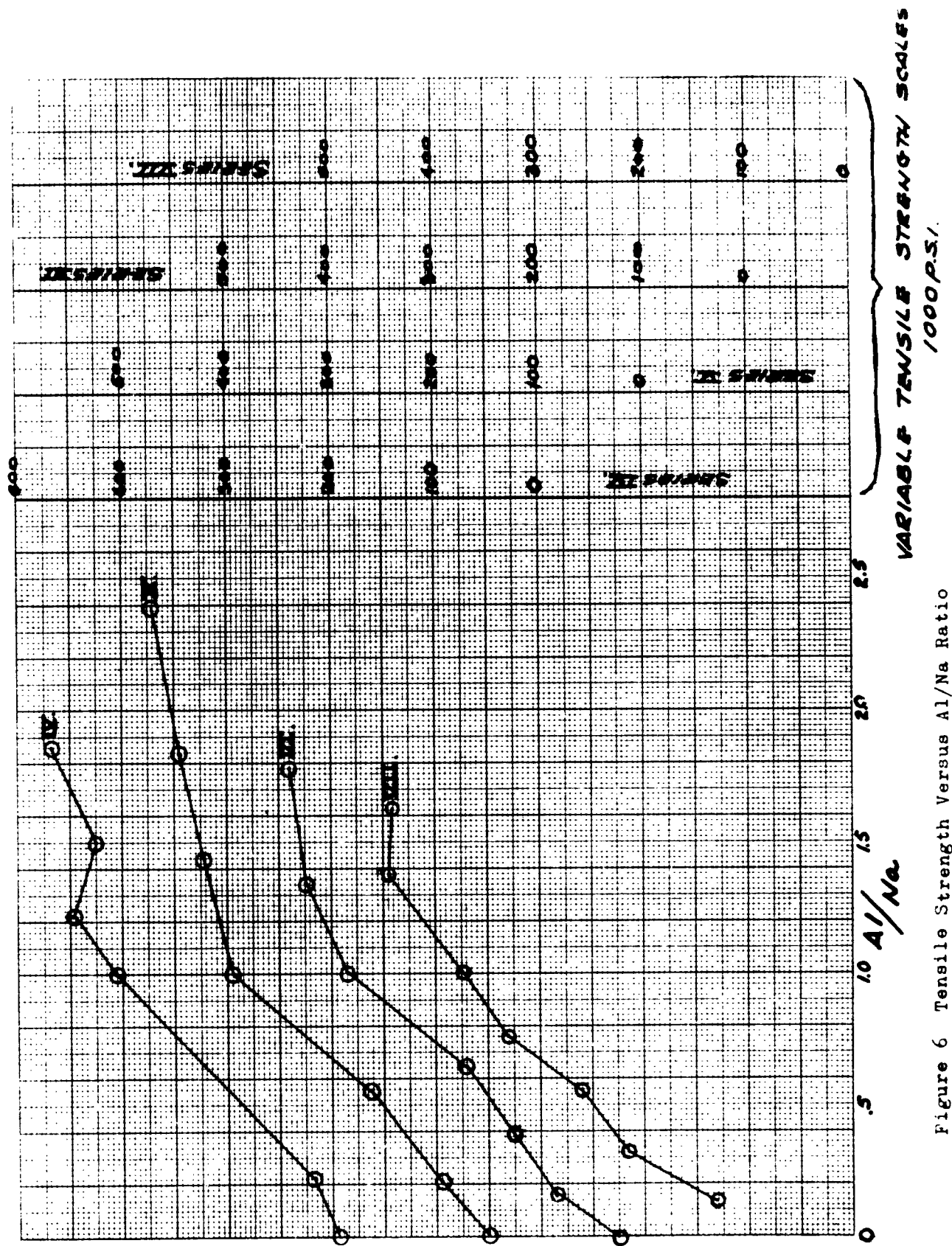


Figure 6 Tensile Strength Versus Al/Na Ratio

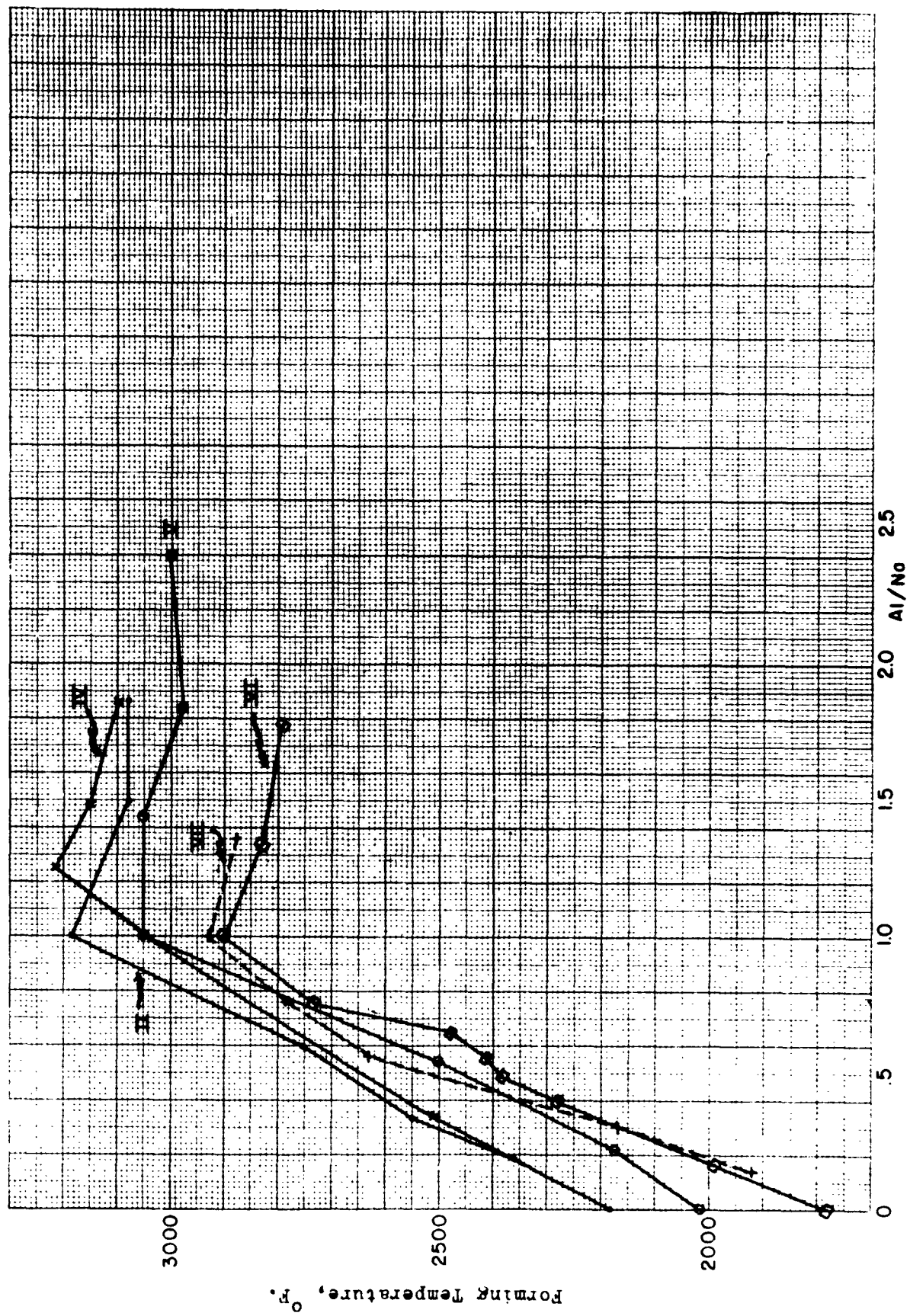


Figure 7 Forming Temperature Versus Al/Na Ratio

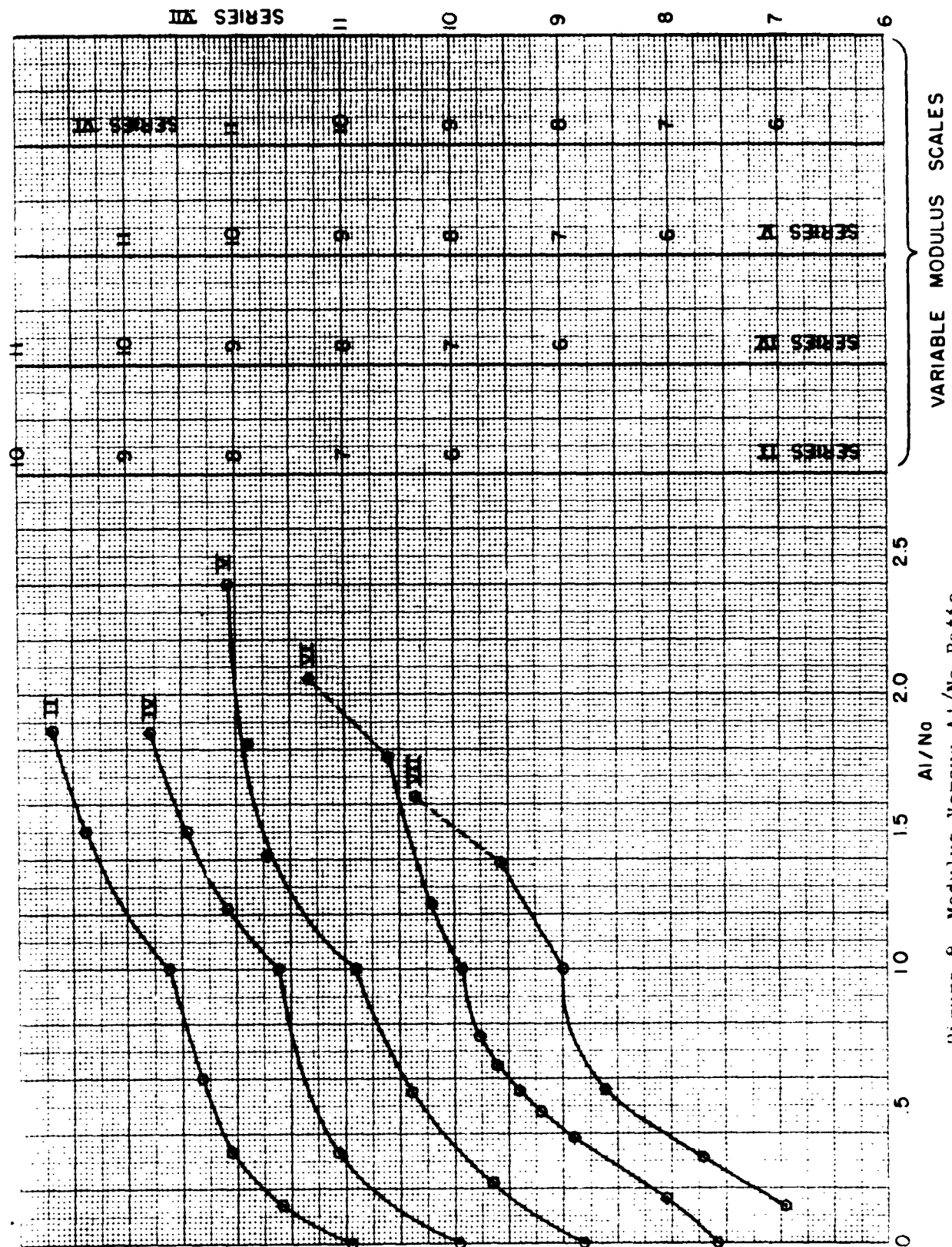


Figure 6 Modulus Versus Al/Na Ratio

5. Experimental Results

a. SiO₂-Al₂O₃-Na₂O System

(1) Tensile Strength

Tensile strength results are summarized in graph form in Figure 6. A separate scale was used for each series of glasses, as indicated on the right of the graph, so that all data could be presented on one sheet without having curves overlap. Tensile strength increases in all cases as the Al to Na ratio increases. It appears that there is an inflection in all curves at or near an Al to Na ratio of one, such that the rate of increase of strength is less at the higher ratios.

(2) Viscosity Measurements

In Figure 7 are given the data for viscosity (fiber forming temperature) vs. the Al to Na ratios of the glasses.

Viscosity increases rapidly with increasing Al to Na ratio up to a ratio of one. Above this, the viscosity levels off or decreases slightly.

(3) Modulus of Elasticity

The data on modulus measurements are plotted in Figure 8. Again a separate scale is used for each series of glasses, so that all curves may be plotted on the same sheet without overlapping. There is an inflection point for all curves at an Al to Na ratio of one. All curves have approximately the same shape. Curves for Series VI and VII may be exceptions, since the locations of the last points indicate that they may rise quite steeply at the higher Al to Na ratios. However, the glasses for which these points were obtained were fiberized with difficulty because of devitrification. These points are therefore questionable. The last parts of the curves are dotted to indicate this.

The data were plotted in various ways to find whether some other significant relationship might be found. Figure 9 shows modulus plotted vs. the ratio of (Si + Al) to O. These curves show an inflection where the ratio is 0.50. This corresponds to the compositions for which the Al to Na ratio is one. However, the curves in Figure 9 show another inflection, where the ratio of (Si + Al) to O is 0.46 or 0.47. Several additional melts were made in Series VI to establish more accurately the location of the inflection and to determine how sharp the break might be. For this series, the inflection appears to be at 0.475 and to follow a smooth curve rather than to have a sharp break. The data are not complete enough or precise enough to show the exact location and nature of the inflections in the other curves. However, all seem to have some kind of a break in about the same region. When plotted in the same way, tensile

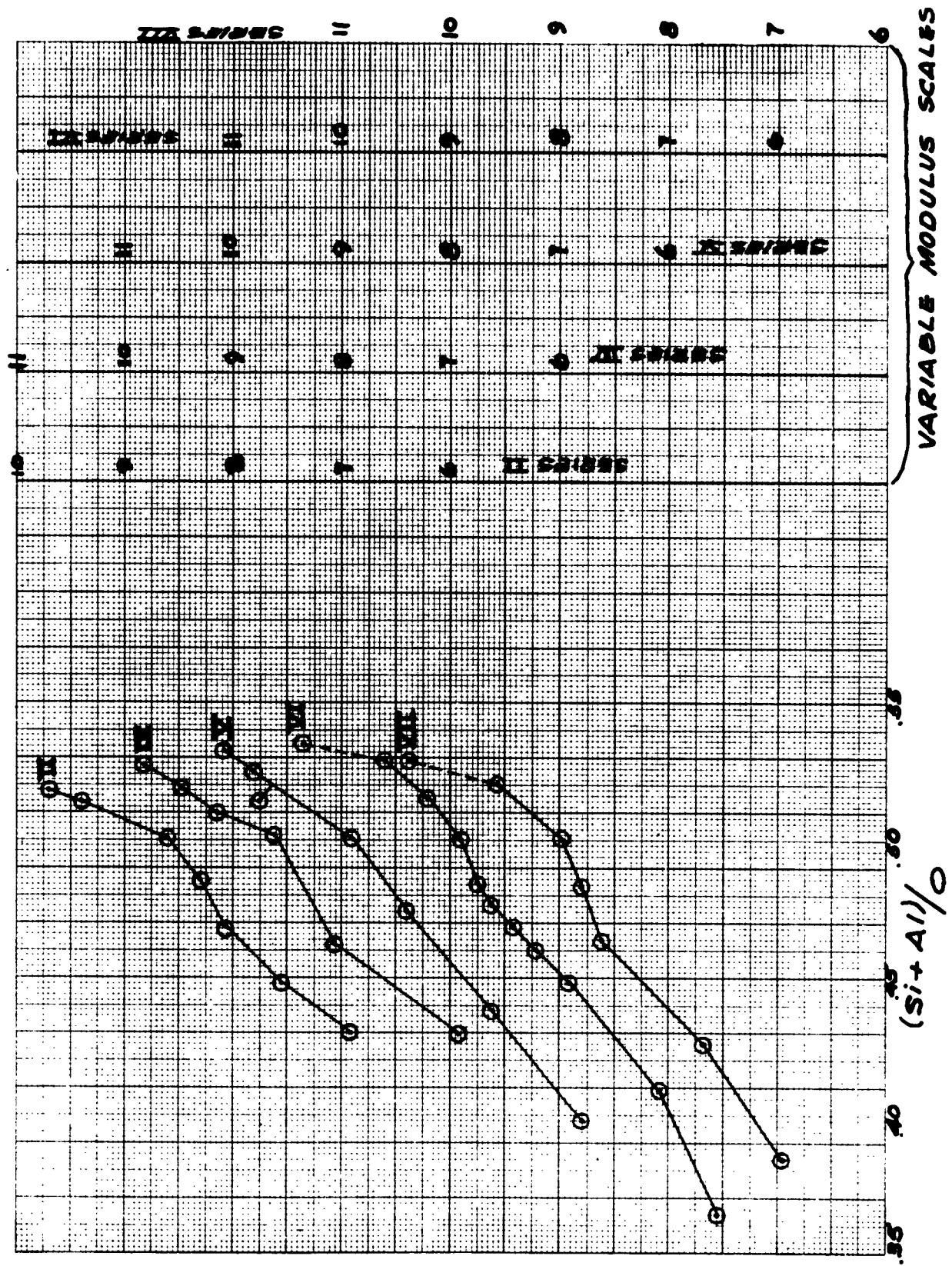


Figure 9 Modulus Versus $(Si + Al)/O$ Ratio

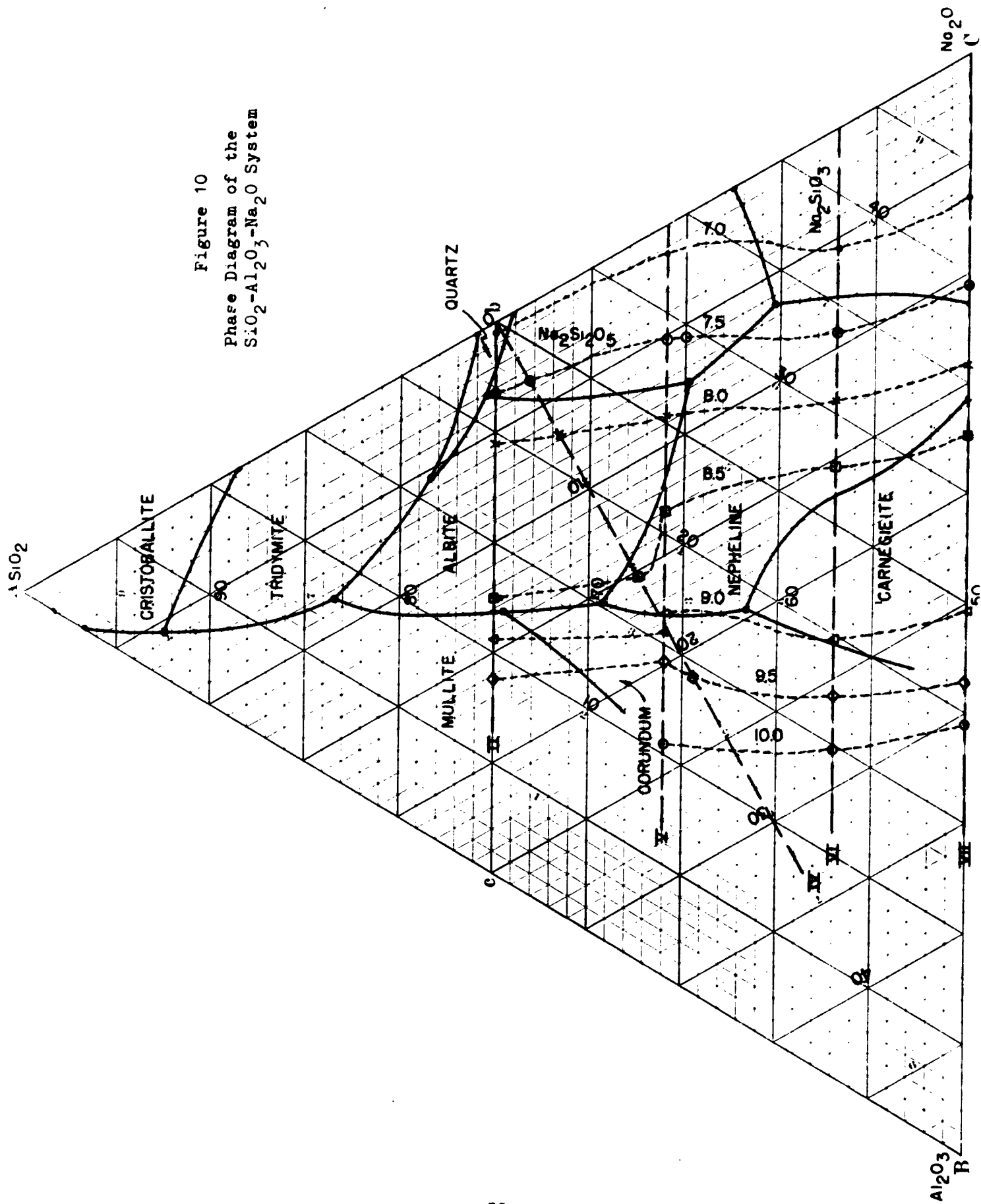
TABLE IV
MODULUS OF AS DRAWN AND ANNEALED FIBERS
SiO₂-Al₂O₃-Na₂O SYSTEM
PSI x 10⁶

<u>GLASS NO.</u>	<u>AS DRAWN</u>	<u>ANNEALED</u>
X-1580A	6.91	8.05
X-1591B	8.06	8.38
X-1592B	8.60	9.60
X-1593A	9.13	10.01
X-1594A	9.44	10.32
X-1595A	9.78	11.01
X-1866	6.79	7.71
X-1867	7.60	8.80
X-1868	8.38	9.51
X-1869	8.87	9.78
X-1870	9.75	10.41
X-1871	9.82	10.60
X-1877	6.53	7.63
X-1878	7.05	8.25
X-1879	7.92	8.99

strength and viscosity curves may also show similar behavior, but results are less well defined. The significance of the breaks in this region has not yet been established.

A number of modulus measurements were made on annealed fibers. Work on annealed fibers was dropped when it became apparent that the information obtained did not justify the additional effort. In general, annealed fibers had moduli about 1,000,000 psi higher than those of fibers as drawn. Table IV lists moduli obtained on fibers as drawn and annealed for those glasses for which data are available.

Modulus data were plotted in a manner such that interpolated values could be obtained for constructing a diagram showing lines of constant modulus vs. composition. Results are shown in Figure 10. The dotted lines are lines of constant modulus of the indicated values. The five series of glasses studied fall on the dashed lines labelled with the appropriate Roman numerals. (Only the top portion of the three-component system from SiO₂ = 50% upwards is shown.) The phase boundaries and primary phases for this system reported in the literature⁽⁴⁾ are also shown. This was done to show whether transitions from regions characteristic of one phase to another would be accompanied by breaks in the constant modulus curves. The data do not show any clear indication of such behavior. The change of modulus from 8.0 to 9.0 x 10⁶ psi requires a larger change in composition than a change of



similar magnitude at either higher or lower values of modulus. The constant modulus curve for modulus of 8.5×10^6 psi shows two breaks, whereas all other lines are nearly straight with only slight curvatures. The region in which this break occurs corresponds to the compositions for which the (Si + Al) to O ratio is 0.46 to 0.47.

This is the range of values in which a second break in the modulus curves was shown in Figure 9.

b. SiO₂-TiO₂-K₂O System

(1) Tensile Strength

Tensile strength results are summarized in graph form in Figure 11. Results are plotted vs. the TiO₂ to K₂O ratio. Any change in properties resulting from a change in the coordination number of the titanium might be expected to be a function of this ratio, since it is related to the amount of oxygen available in the system for coordinating with the titanium.

At the 80% SiO₂ level, it appears that there is a break in the curve at a ratio of one. At the 70% level, the curve is smooth, but it does have an inflection in this region. At the lower SiO₂ levels, it is not possible to be certain whether or not there is any meaningful break in the curves. The precision of the measurements is lower at the very low strengths involved. In addition, the glasses with the lower ratios have such poor durability that chemical attack on the fibers by the moisture of the air during the testing time could contribute to inaccuracies.

Examination of the published data on other physical properties of glasses in this system suggested that the position at which properties changed was dependent also upon the SiO₂ content. For this reason, the curves shown in Figure 12 were prepared showing the tensile strength as a function of the ratio of SiO₂ + TiO₂ to K₂O. The curves for 70 and 80% SiO₂ glasses are approximately parallel, but the other two do not appear to fit this pattern.

(2) Viscosity Measurements

The fiber forming temperatures are plotted as a function of TiO₂ to K₂O ratio in Figure 13. There is a suggestion of a break in the curve for the 80% SiO₂ series at a ratio of somewhat less than one. The rest of the curves approach a straight line relationship very closely.

Two curves are given for the 60% SiO₂ series. The one drawn as a solid line was obtained when the fibers tested for strength and modulus were run. When it became apparent that these results were inconsistent, the glasses were re-run to

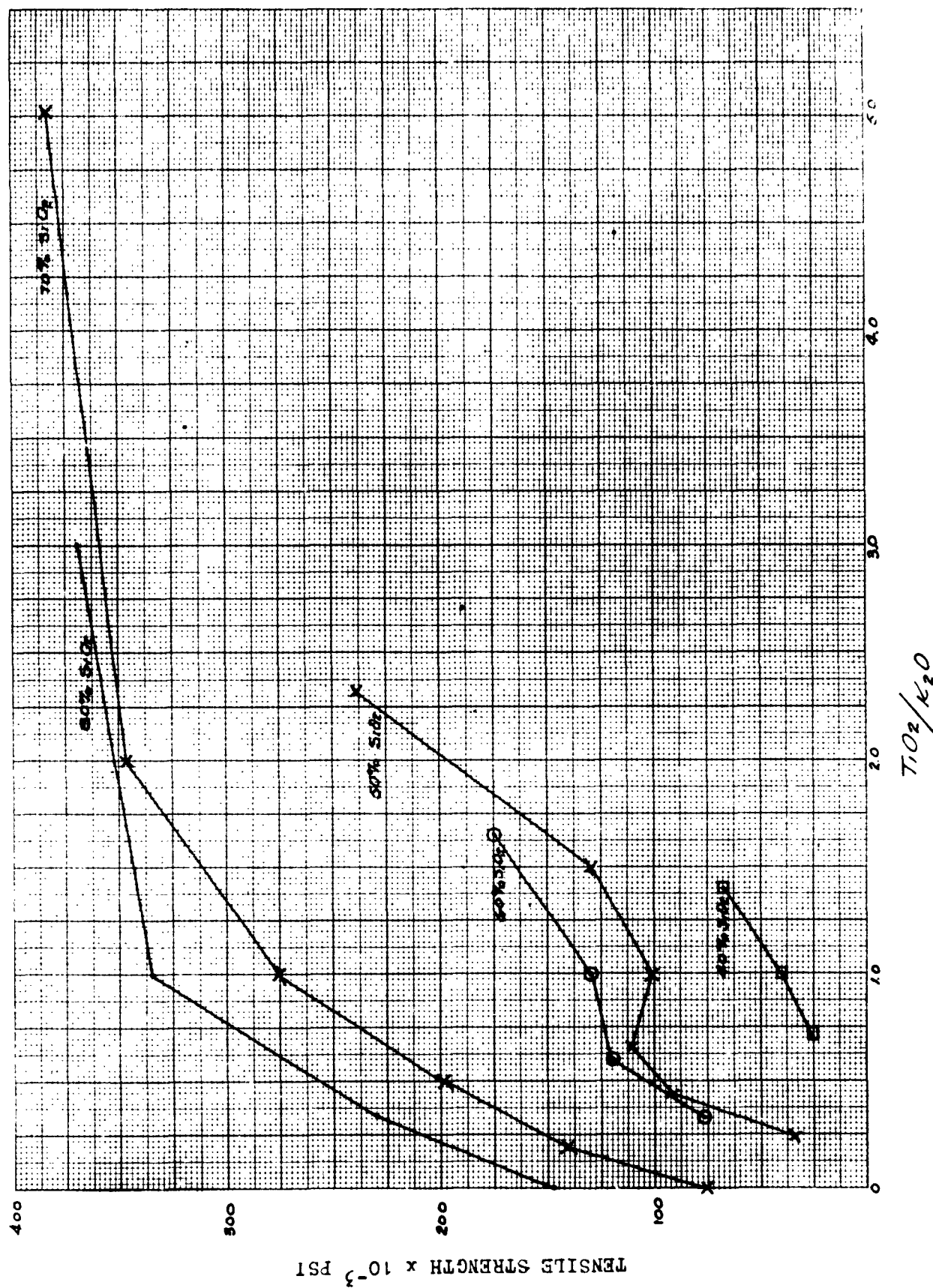


Figure 11 Tensile Strength Versus TiO₂/K₂O Ratio

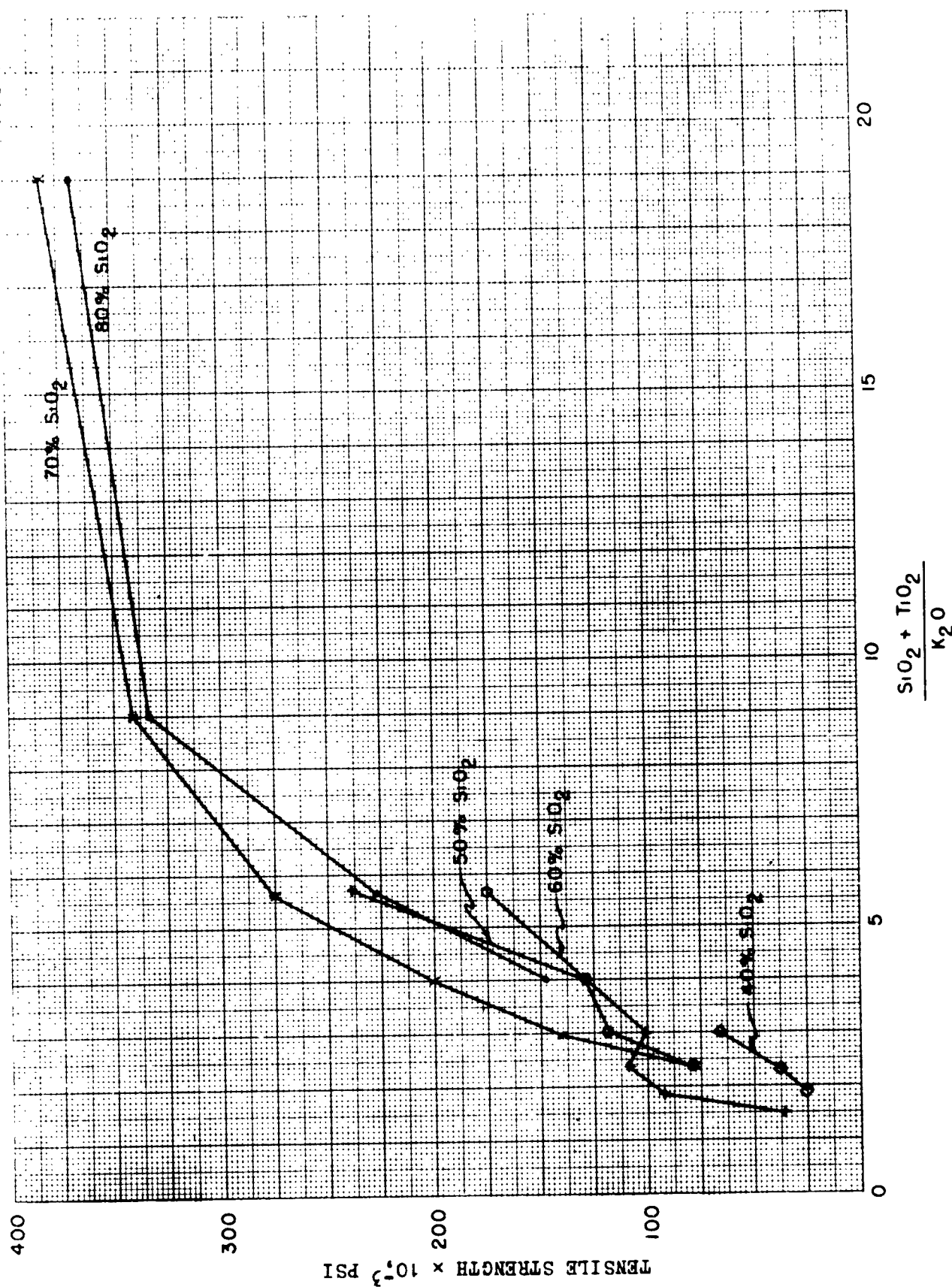


Figure 12 Tensile Strength Versus $(\text{SiO}_2 + \text{TiO}_2)/\text{K}_2\text{O}$ Ratio

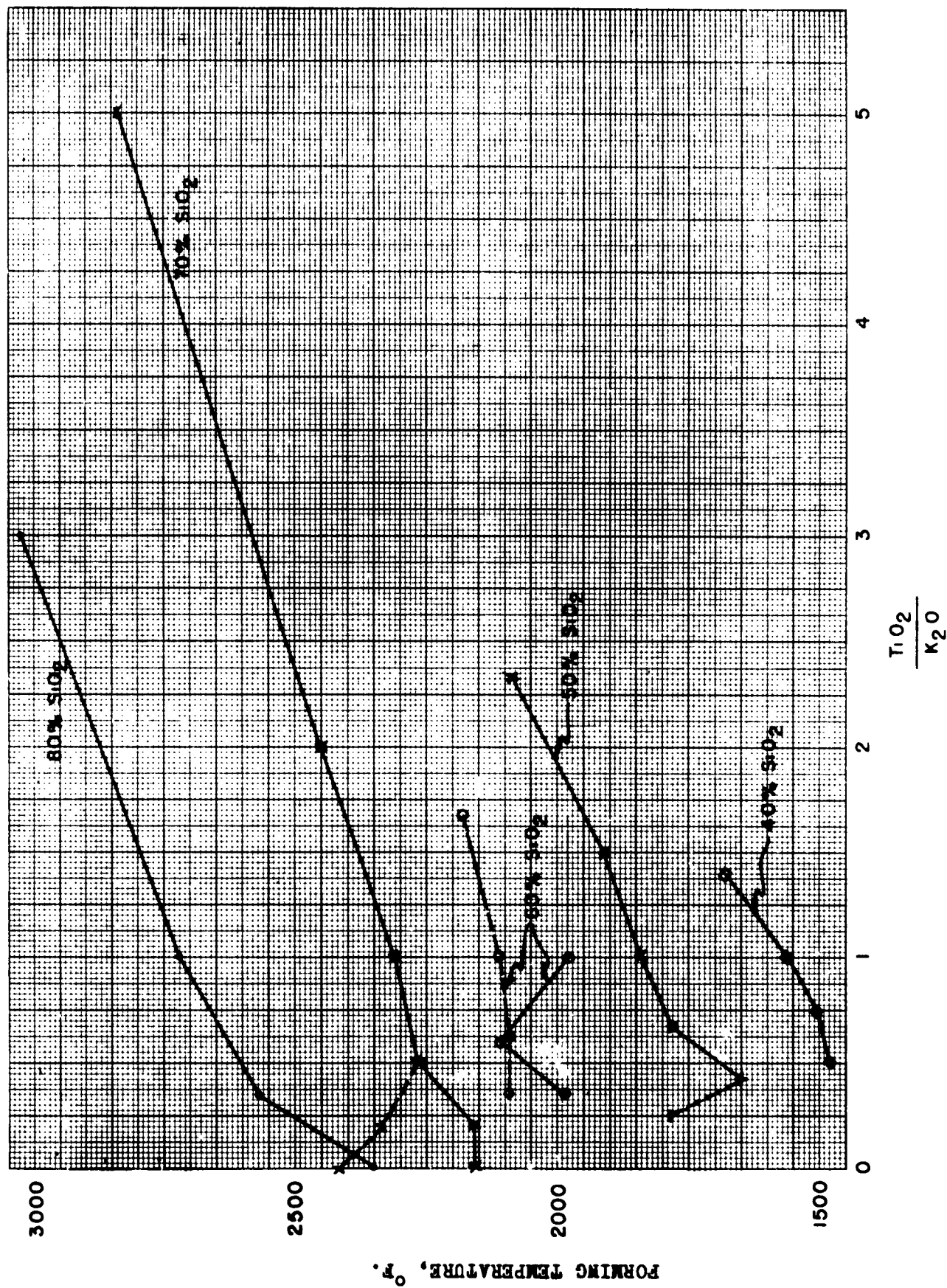


Figure 13 Forming Temperature Versus $\text{TiO}_2/\text{K}_2\text{O}$

determine fiber forming temperatures only. The results are shown on the curve drawn in a dotted line. The re-run indicated that these glasses must have a very flat viscosity curve, that is, a relatively large change in temperature is required to produce a given change in fiber diameter. This could account in part for the unusual shape of the original curve. Care was taken to obtain forming temperatures at very nearly the same fiber diameters on the second run. Therefore, this curve is to be regarded as more reliable. The second curve was obtained with a different bushing. Small differences in bushing construction or thermocouple output may have caused the position of the curve to be generally higher than that obtained on the first run. The slope of the curve seems to be less at the 60% SiO_2 level than at either higher or lower levels.

The point at the lowest TiO_2 to K_2O ratio for the 50% SiO_2 series was obtained the day the thermocouple on the bushing failed. Its accuracy is therefore suspect.

Forming temperature data were re-plotted in Figure 14, showing forming temperature vs. the $\text{SiO}_2 + \text{TiO}_2$ to K_2O ratio. The curves have about the same shape and relationship to each other as shown on Figure 13.

(3) Modulus of Elasticity

Modulus results are plotted as a function of the TiO_2 to K_2O ratio in Figure 15. The curves are all similar in shape. The curve at the 70% SiO_2 level appears to be out of place at the upper end in that it crosses those for the 50% and 60% SiO_2 levels. The other curves seem to approach the same value at a TiO_2 to K_2O ratio of 3.0.

None of the curves show a sharp break, although all curve downward in the general vicinity where the TiO_2 to K_2O ratio is one.

When modulus is plotted vs. the $\text{SiO}_2 + \text{TiO}_2$ to K_2O ratio, Figure 16, all curves seem to fall into a consistent pattern. One point on the curve for the 50% SiO_2 level appears to be off, but it did not fit the apparent curve in Figure 15 either.

4. Discussion of Results

In the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-Na}_2\text{O}$ system, the three properties measured- fiber forming temperature, tensile strength, and modulus of elasticity - appear to be dependent upon glass structure in the same manner as other physical properties reported in the literature. A break in composition vs. property curves was found where the Al to Na ratio for the glass was one. The suggestion that this is related to a change in coordination number of the alumina is consistent with these observations. Glass viscosity (as shown by fiber forming temperature)

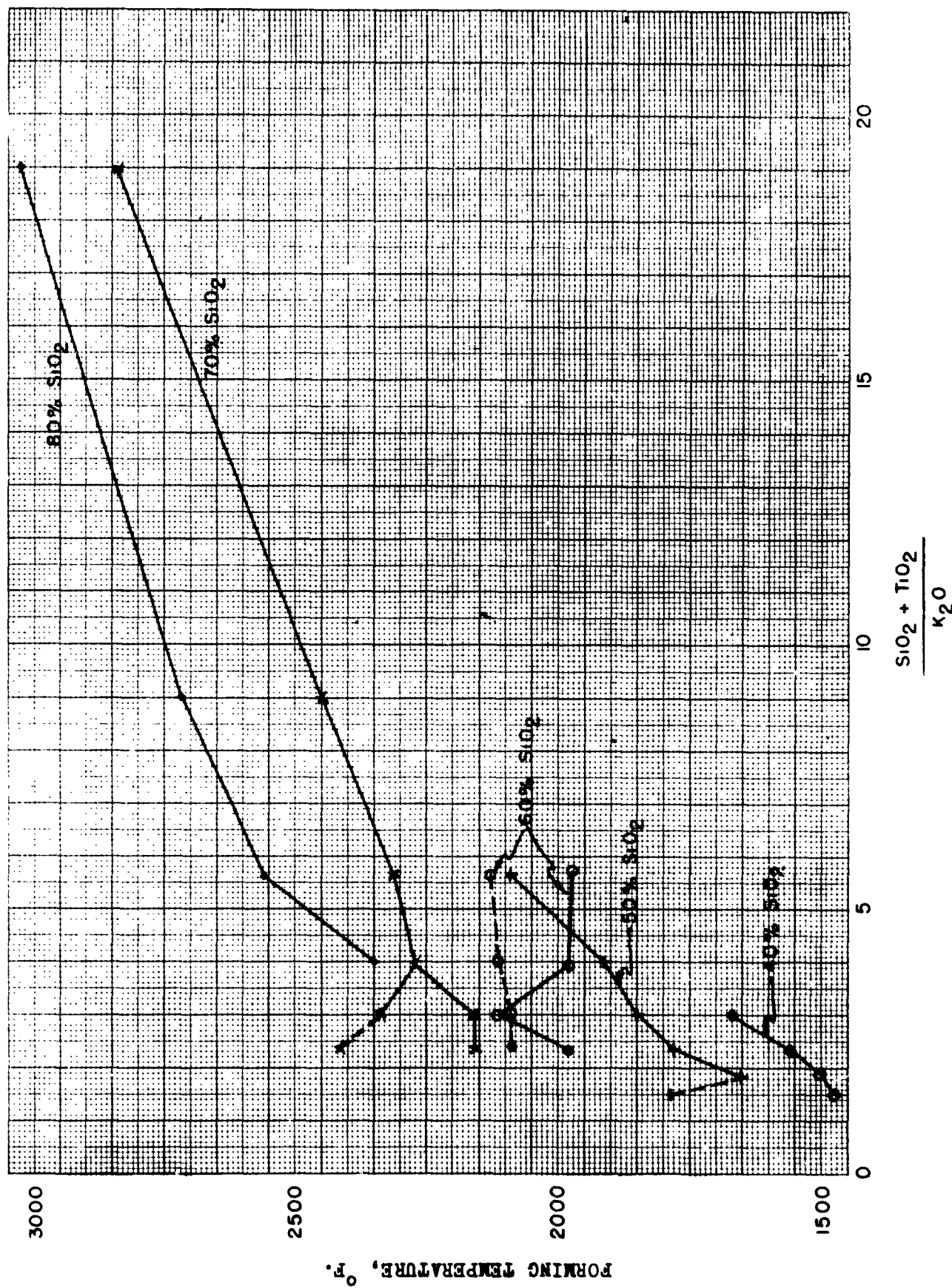


Figure 14 Forming Temperature Versus $(\text{SiO}_2 + \text{TiO}_2) / \text{K}_2\text{O}$ Ratio

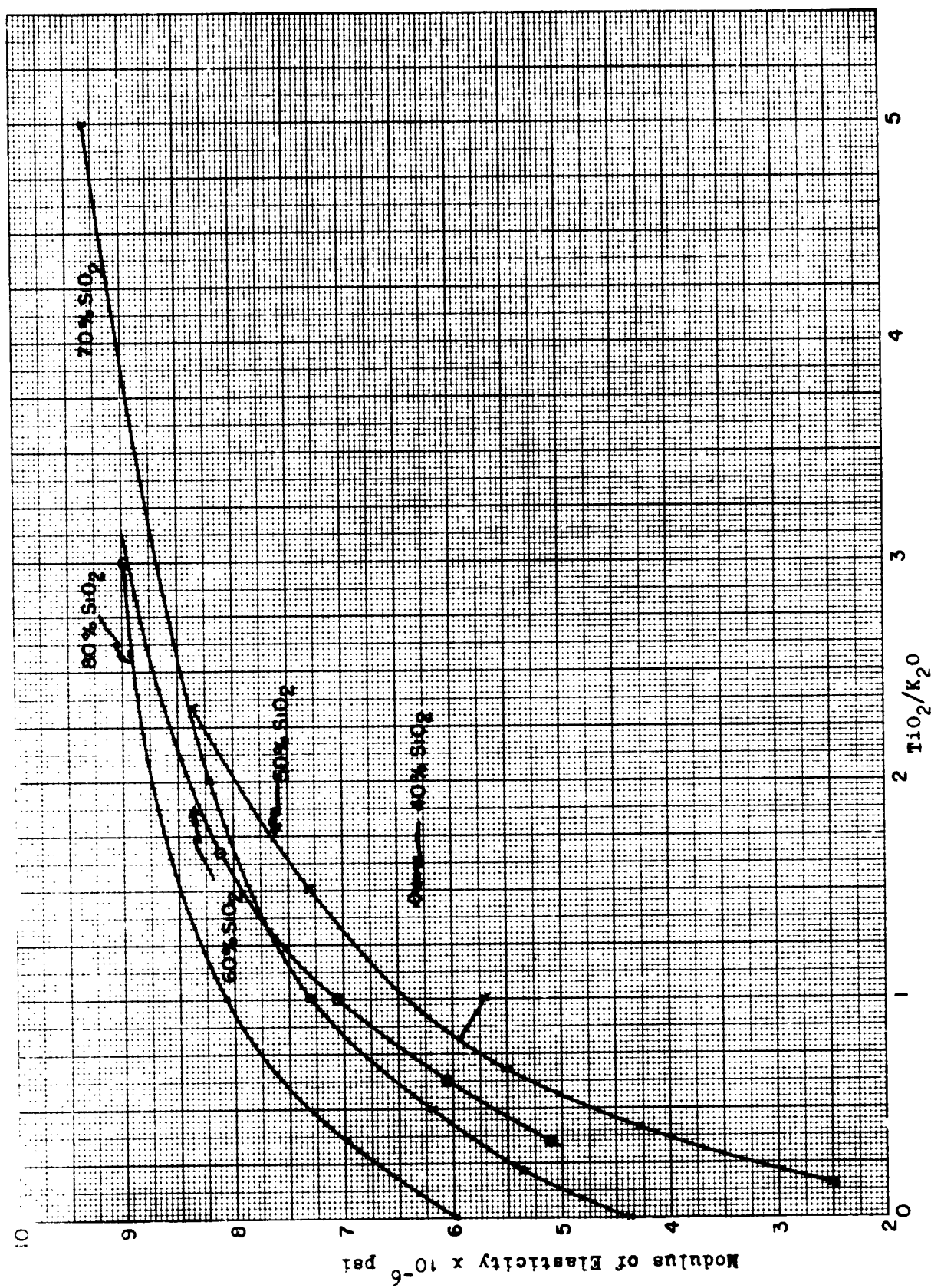


Figure 15 Modulus Versus $\text{TiO}_2/\text{K}_2\text{O}$ Ratio

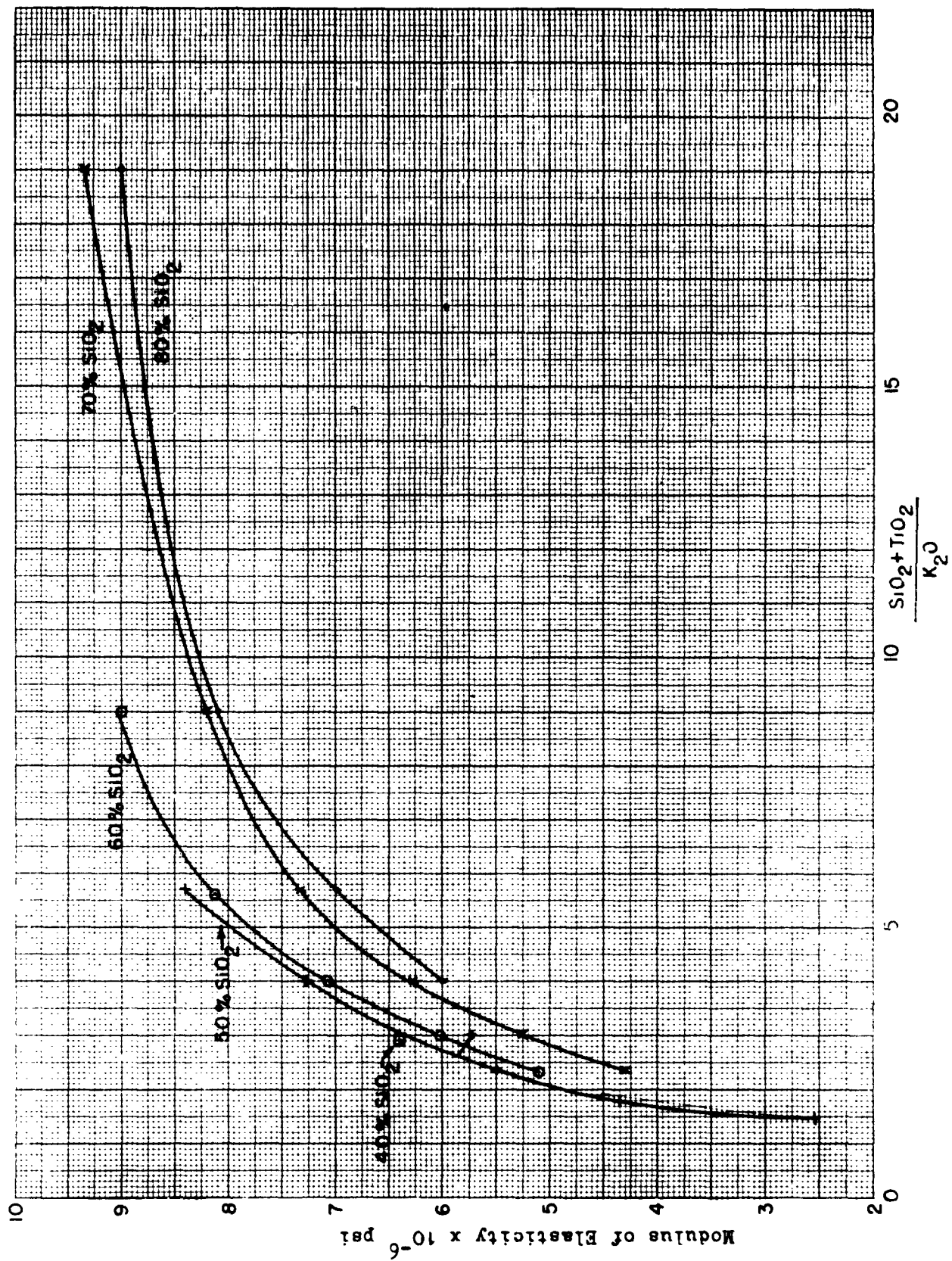


Figure 16 Modulus Versus $(\text{SiO}_2 + \text{TiO}_2) / \text{K}_2\text{O}$

increases at first as Al_2O_3 is substituted for Na_2O . This is logical if the Al_2O_3 enters the network and behaves much as SiO_2 would. At higher Al_2O_3 contents where the Al to Na ratio exceeds one, the aluminum added enters the structure with a coordination number of six, and acts as most glass modifiers do, by lowering the viscosity. Similarly, substitution of Al_2O_3 for Na_2O increases the tensile strength, by increasing the proportion of ions participating in the network. Strength continues to increase, but at a lower rate, at ratios of Al to Na higher than one. It could be expected that viscosity might be related to tensile strength in that both are dependent upon the forces holding the network together, and that if a change in composition caused the viscosity to drop, the tensile strength should drop also. Two possible reasons for the results actually obtained are suggested:

- a. Forming temperature may not be a simple function of viscosity. Surface tension, for instance, can affect forming behavior, and might change in such a way as to counteract changes in viscosity.
- b. The cooling of the glass may be accompanied by further structural changes which make it behave differently at room temperature than it would at elevated temperatures.

Measurements of the modulus of elasticity show the break at the Al to Na ratio of one more clearly and consistently than to tensile strength or forming temperature. This is not surprising. Tensile strength of fibers is known to be closely related to the conditions under which they were formed. Attempts were made to keep forming conditions similar for the fibers tested in this study. However, it is quite possible that if enough conditions had been tried to find the optimum strength for each composition, some inconsistencies in the data would have disappeared. The measurement of forming temperature was just incidental, and was not the primary purpose of the work. Hence, precautions that could have been taken to increase accuracy were not observed, and results are expected to contain small errors. Modulus results, on the other hand, seem to be less sensitive to variations in forming conditions. The precision is believed to be relatively good, at least when the diameters of the fibers tested lie in a small range. The spread in results for the eight or more determinations made for each point reported was usually not more than 0.3×10^6 psi, and normally less than this.

The nature of the change in modulus as the Al to Na ratio increases is somewhat surprising. A structural change which results both in lowered viscosity and a reduction in slope of the tensile strength curve causes an increase in the modulus. It may be that the larger size structural units resulting from six-coordinated aluminum can increase the rigidity of the structure without increasing its strength.

Recently Lacy⁽⁵⁾ offered another proposal for the structure of glasses in the SiO_2 - Al_2O_3 - Na_2O system. He suggested that the logical packing

of ions favors the formation of what he calls "triclusters." These triclusters consist of three tetrahedral groups having a common oxygen, which is shared by an AlO_4 group and two SiO_4 groups or an SiO_4 and two AlO_4 groups. Each two triclusters conserve one oxygen. When the Al to Na ratio is greater than one, there are not enough oxygens for all glass formers to be in tetrahedral groups. According to Lacy, the formation of triclusters compensates for this oxygen shortage. This theory would probably account for the observed results as well as the concept of change in coordination numbers if we assign suitable properties to the triclusters. The data do not seem to support one concept over the other.

The properties measured on the glasses in the SiO_2 - TiO_2 - K_2O system do not appear to lead to as good a concept of glass structure. Examination of Figures 15 and 16 shows that the modulus increases as either the TiO_2 to K_2O or the $\text{SiO}_2 + \text{TiO}_2$ to K_2O ratio increases. The rate of increase drops off at the higher ratios. The absence of any sharp break in the curves suggests that the change in slope does not result from a sudden change in structure at some particular composition. If, for instance, a change in coordination number of the titanium is responsible, it seems likely that titanium is present to some degree in both states in all glasses studied. Changes in composition thus probably change the relative amounts in some rather gradual fashion. That the change in structure is affected also by the SiO_2 content is suggested by the fact that the data seem to fall into a systematic system of curves in Figure 16, while there is some crossing of curves in Figure 15, in which no account is made of the SiO_2 content.

It might be argued that Figures 11 and 12 do not support the latter contention in that curves for tensile strength seem to show a more orderly relationship when plotted vs. the TiO_2 to K_2O ratio than when plotted vs. the $\text{SiO}_2 + \text{TiO}_2$ to K_2O ratio. However, examination of Figure 12 suggests that it is possible that the proper order of the curves for the various series of glasses might be the same as that found for modulus. At the very low ratios of $\text{SiO}_2 + \text{TiO}_2$ to K_2O , the curves tend to fall in the same order as the modulus curves in Figure 16. The peculiar shapes of the tensile strength curves at the 50% and 60% SiO_2 levels do not seem logical and may indicate experimental error.

Figure 17 was plotted to assist in plotting lines of constant modulus on a three-component diagram. It appears that at any SiO_2 level the modulus may be a linear function of the K_2O content of the glass.

The constant modulus curves are plotted in Figure 18. These curves appear to offer no particularly significant implications relative to glass structure.

The SiO_2 - Al_2O_3 - K_2O system should be studied as the SiO_2 - Al_2O_3 - Na_2O and the SiO_2 - TiO_2 - K_2O systems have been. Results should be studied to determine if some correlation between the two systems can be found.

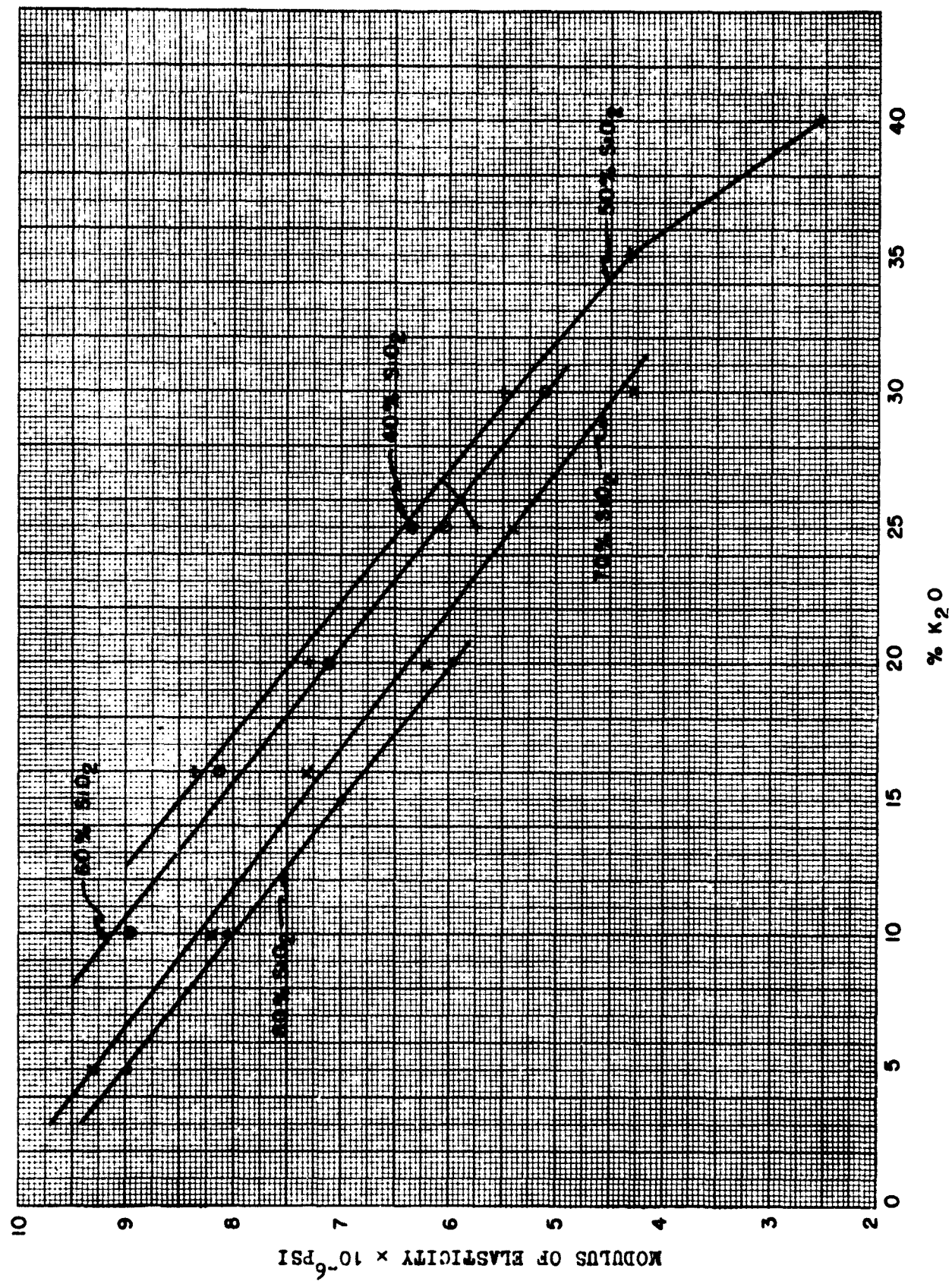
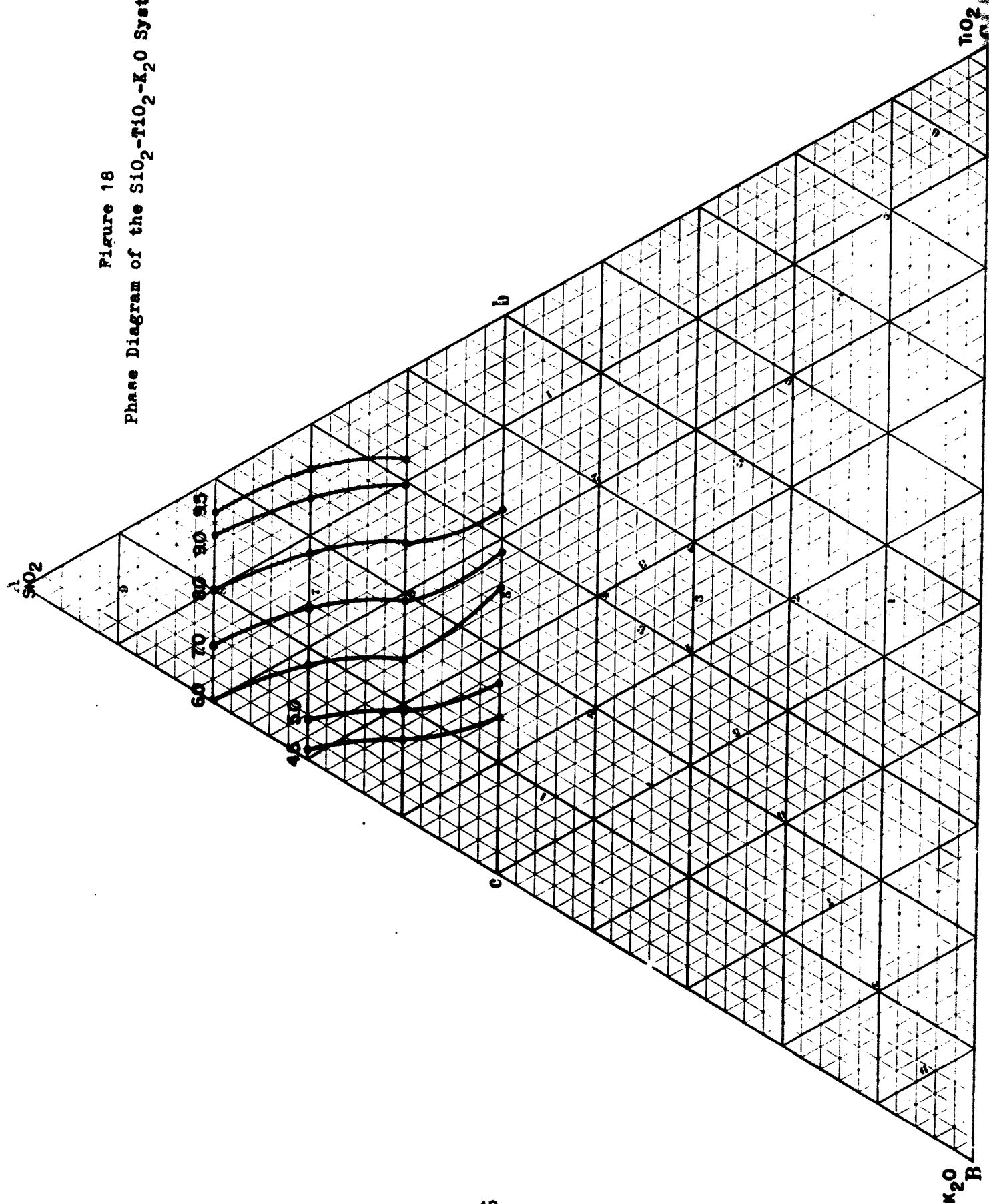


Figure 17 Modulus Versus Percent K₂O

Figure 18
Phase Diagram of the SiO_2 - TiO_2 - K_2O System



Differences between the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-K}_2\text{O}$ system and the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-Na}_2\text{O}$ system should be examined to see if they can be related to the size of the alkali ion added. If such relation is found, study of a system utilizing a third alkali ion, probably lithium, should be undertaken.

III CONCLUSIONS

The results of this investigation indicate that glass fibers possessing the highest tensile strength and modulus of elasticity values are located in the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-MgO}$ and the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-ZnO}$ systems. There is some indication that the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-BeO}$ system may also possess desirable properties.

The best glass compositions in the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-MgO}$ system fall at an Al_2O_3 to MgO mol ratio of approximately 1.0 and contain from 60 to 70% by weight of SiO_2 . The $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-ZnO}$ system has not been investigated extensively, but it is expected that the glasses with the best properties are located within a similar range of values. Investigation of the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-BeO}$ system should be continued, but is complicated by the extreme devitrification tendencies of this glass.

The addition of a fourth oxide component to a base glass consisting of the S-994 composition does not improve the tensile strength or modulus of elasticity of the resultant glass. The addition of minute quantities of a fourth oxide component has received only preliminary study, but might prove beneficial.

The exact effect of fiber diameter on both tensile strength and modulus is still open to question. However, there is some evidence that a reduction in diameter of $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-MgO}$ and $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-ZnO}$ glass fibers does improve properties, especially the modulus of elasticity.

IV
REFERENCES

- (1) D.E. Day and G.E. Rindone. Properties of Soda Aluminosilicate Glasses: I. Refractive Index, Density, Molar Refractivity and Infrared Absorption Spectra. J. Am. Ceram. Soc., V. 45, No. 10, p. 489-496 (1962)
- (2) D.E. Day and G.E. Rindone. Properties of Soda Aluminosilicate Glasses: II Internal Friction, Ibid, p. 496-504.
- (3) Bh. V. Janakirama Rao. The Dual Role of Titanium in the System $K_2O - SiO_2 - TiO_2$. Phys. and Chem. of Glasses, V.4, #1, 1963, p. 22-34
- (4) E.M. Levin, C.R. Robbins and H.F. McMurdie . Phase Diagrams for Ceramists, The American Ceramic Society, 1964.
- (5) E.D. Lacy. Aluminum in Glasses and in Melts, Phys. and Chem. of Glasses, V. 4, # 6, 1963, p. 234-238.

DISTRIBUTION LIST

Commander
Attn: HDLR
Holloman AFB, New Mexico

RTD (RTHM, Maj. R. Hemm)
Bolling AFB, D.C. 20332

RTD (RTHM, Dr. S.W. Strauss)
Bolling AFB, D.C. 20332

AFSWC (Technical Library)
Kirtland AFB, New Mexico

Commander
Attn: E.L. Cole, Lt/Col USAF
Director, Bioastronautics Project
Air Force Ballistic Missile Division
Air Force Unit Postoffice
Los Angeles, California, 90045

SSD (SSTRE, Maj. Iller)
AF Unit Postoffice
Los Angeles, California, 90045

Director, Air University
Attn: 7575
Maxwell AFB, Alabama

MCSA
Attn: Mr. Norman Sachs
Philadelphia 46, Pennsylvania

Sandia Corporation
Attn: Mr. A. Pope
Sandia Base
Albuquerque, New Mexico

RTD (RTT-Col. J.V. Hearn, Jr.)
Bolling AFB, D.C. 20332

Hq. USAF (AFRSTG, Lt. Col. H. Hamlin)
Washington D.C. 20330

AFSC Scientific & Technical Liaison Office
Attn: Mr. John J. Custer, Staff Engineer
111 East 16th Street
New York 3, N.Y.

Department of the Army
Army Research Office
Office of Chief of Research & Development
Attn: Dr. B.S. Fisher
Washington 25, D.C.

Hq. Department of the Army
Attn: Lt/Col. L.G. Klinker
Office, Chief of Research & Devel.
Army Research Office
Washington 25, D.C.

Army Research Office
Attn: (Mr. Allan L. Tarr, Tech. Sec'y.)
Office, Chief of Research & Devel.
3045 Columbia Pike
Arlington 4, Virginia

U.S. Army Engineer Research &
Development Laboratories
Attn: Mr. Arthur W. Van Heuckeroth
Fort Belvoir, Virginia

Hq. Signal Research & Development
Laboratories
Attn: Dr. S. Benedict Levin
Institute for Exploratory Research
Fort Monmouth, New Jersey

Hq. Signal R & D Laboratory
Attn: Mr. Louis Reiss
Fort Monmouth, New Jersey

Office, Chief of Ordnance
Department of the Army
Attn: ORDTB, Materials
(Mr. T.B. Blevins)
Washington 25, D.C.

Army Ballistic Missile Agency
(ORDAB-DS)
Redstone Arsenal
Huntsville, Alabama

Ordnance Materials Research Office
Attn: Mr. George A. Darcy, Jr.
Watertown Arsenal
Watertown 72, Massachusetts

U.S. Army
Natick Laboratories
C & O Materials Division
Attn: Dr. S.J. Kennedy
Natick, Massachusetts

U.S. Army
Natick Laboratories
C & O Materials Division
Attn: Dr. George R. Thomas
Natick, Massachusetts

Office of Advanced Research Programs
Attn: Mr. Norman Mayer
NASA
Washington 25, D.C.

NASA
Langley Research Center
THRU: AF Development Field Office
Langley AFB, Virginia

NASA
Research Materials Division
Attn: Mr. Bernard Achhammer
1520 H. Street, N.W.
Washington 25, D.C.

NASA
Lewis Research Center
Attn: Mr. John W. Weeton, Chief
Composite Materials Branch
2100 Brookpark Road
Cleveland 35, Ohio

Chief
Bureau of Naval Weapons
Attn: AE-5
Washington 25, D.C.

Bureau of Naval Weapons
Attn: Mr. N.E. Promisel, R-15-Rm 2W98
Washington 25, D.C.

U.S. Naval Ordnance Laboratory
Attn: Dr. Albert Lightbody, Chief
Code W, Ext. 672
White Oak
Silver Spring, Maryland

U.S. Naval Laboratory
Attn: Mr. Gordon L. Fogal
White Oak
Silver Spring, Maryland

DDC : THRU
AFML (MANF), Mr. W.H. Gloor
Wright-Patterson AFB, Ohio 45433

Aerojet General Corporation
Attn: Mr. A. Lewis
Structural Materials Division
P.O. Box 296
Azusa, California

Aerospace Corporation
Attn: Mr. John C. Batteiger
SR & P Division
2400 El Segundo Blvd.
El Segundo, California

Air Research Manufacturing Co.
Attn: Mr. V.K. Crawford
Supervisor of Materials Laboratory
9851 Sepulveda Blvd.
Los Angeles 45, California

Allied Chemical Corporation
Attn: Dr. Hans R. Naumark
Assistant to the Director
Research & Development
P.O. Box 405
Morristown, New Jersey

Argonne National Laboratory
Attn: Mr. Hoylande D. Young
Technical Information Bureau
9700 South Cass Avenue
Argonne, Illinois

ASTRO Research Corporation
Attn: Mr. H.U. Schuerch
1330 Cacique Street
Santa Barbara, California

AVCO Research & Development Corp.
Attn: Mr. J.W. Herrick
Materials Dept., Plastics Section
201 Lowell Street
Wilmington, Massachusetts

Babcock & Wilcox Company
Attn: Mr. F.J. Hartwig
Research & Development Division
Research Center
Alliance, Ohio

Battelle Memorial Institute
Attn: Mr. Robert Leininger
Polymer Research-Chemistry Department
505 King Avenue
Columbus 1, Ohio

Beech Aircraft Corporation
Attn: Mr. Chester A. Rembleske, Chief
Administrative Engineer
Wichita 1, Kansas

Bell Aerosystems Company
Attn: Mr. George Kappelt, Director
Engineering Laboratories
Buffalo 5, New York

Bell Helicopter Company
Attn: Mr. George Linnabery
Chief Development Engineer
Fort Worth 1, Texas

Bell Telephone Laboratories, Inc.
Attn: Mr. T.V. Wallder, Head
Plastics Development Department
Murray Hill, New York

Bendix Products Division
Attn: Mr. Frank C. Albright,
Staff Engineer
Systems Development Section
3300 South Sample Street
South Bend, Indiana

Bjorksten Research Laboratories
P.O. Box 265
Madison 1, Wisconsin

Boeing Airplane Company
Attn: Mr. Cecil B. Barlow, Chief
Design Services
Wichita, Kansas

Boeing Airplane Company
Attn: Mr. George M. Powell, Chief
Materials & Processes Engineer
Vertol Division
Morton, Pennsylvania

Boeing Airplane Company
Attn: Mr. Harold A. Welch, Mgr.
Engrg. Laboratories, Box 27-73
Transport Division
Seattle 34, Washington

The Carborundum Company
Ceramic Fiber Plant
Attn: Mr. Theodore J. Cybularz
Refractories & Electronics Div.
P.O. Box 337
Niagara Falls, New York

Chance Vought Corporation
Attn: Mr. Milton J. Rudick
Chief of Applied R & D
Astronautics Division
P.O. Box 5907
Dallas, Texas

Chemstrand Research Center, Inc.
Attn: Dr. C. Stehman
P.O. Box 731
Durham, North Carolina

Convair
A Div. of General Dynamics Corp.
Attn: Mr. H.A. Swift, Chief
Structures and Materials
Pomona, California

Convair
A Div. of General Dynamics Corp.
Attn: Mr. E.F. Strong, Chief
Engineering Laboratories
Mail Zone 63-10
San Diego 12, California

Convair
A Div. of General Dynamics Corp.
Attn: Mr. J.F. Robinson, Chf. Engr.
Test Laboratories
Fort Worth, Texas

Cook Research Laboratories
Div. of Cook Electric Company
Attn: Mr. Fred A. Ruprecht
Aerospace Technology Section
6401 W. Oakton Street
Morton Grove, Illinois

Corning Glass Works
Research and Development Division
Attn: Mr. J.M. Williams
Corning Glass Works Laboratories
Corning, New York

Douglas Aircraft Company, Inc.
Attn: Mr. E.P. Troeger, Chief
Materials Research & Process Engr.
3000 Ocean Park Blvd.
Santa Monica, California

Douglas Aircraft Company, Inc.
Attn: Mr. Frank M. Wright
Space Structures Section
3000 Ocean Park Blvd.
Santa Monica, California

E.I. duPont de Nemours & Company, Inc.
Attn: Mr. R.W. Fulling
Development Dept. (Special Interests)
duPont Building
Wilmington 98, Delaware

Engelhard Industries, Inc.
Attn: Mr. E.M. Scott
Amersil Quartz
685 Ramsey Avenue
Hillside, New Jersey

Fabric Research Laboratories, Inc.
Attn: Dr. K.R. Fox
V.P. in Charge of Development
1000 Providence Highway
Dedham, Massachusetts

Fiber Glass Industries, Inc.
Attn: Mr. Hamilton Bickford
5 Alman Street
Broadalbin, New York

FMC Corporation
Chemical R & D Center
Attn: Mr. P.E. Willard, Asst. Director
Government Liaison
P.O. Box 8
Princeton, New Jersey

Forest Products Laboratories
Attn: Dr. Richard F. Blomquist
Organic Chemist
Division of Timber Processing
Madison 5, Wisconsin

Fram Corporation
Attn: Mr. C.A. Rodman
Research Department
55 Pawtucket Avenue
East Providence 16, Rhode Island

General American Transportation Corp.
Attn: Mr. R.A. Bambenek
Mechanics Research Division
7501 N. Natchez Avenue
Niles 48, Illinois

General Dynamics/Astronautics
Attn: Mr. R.H. Pearson
Materials Research
5001 Kearney Villa Road
San Diego, California

General Electric Company
Missiles & Space Vehicle Department
Attn: Mr. L. McGreight
3750 "D" Street
Philadelphia 24, Pennsylvania

General Electric Company
Attn: Mr. R.F. Koenig, Mgr.
Materials & Processes Engrg.
MSD Valley Forge STC-ro U8417
P.O. Box 8555
Philadelphia 1, Pennsylvania

General Technologies Corporation
Attn: Mr. George R. Thompson
708 N. West Street
Alexandria, Virginia

Gruman Aircraft Engrg. Corp.
Attn: Mr. A.R. Mead
Manufacturing Staff Engr.
Bethpage, Long Island, N.Y.

The B.F. Goodrich Research Center
Attn: Dr. L. Loughborough
Brecksville, Ohio

Goodyear Aircraft Corporation
Attn: Mr. D.M. Marco
Materials Technology Department
1210 Massillon Road
Akron 15, Ohio

Goodyear Tire & Rubber Company
Attn: Dr. James D. D'Ianni
Research & Development
Akron 16, Ohio

Harris Research Laboratories, Inc.
Attn: Mr. J.F. Krasny
6220 Kansas Avenue, N.E.
Washington 11, D.C.

Hercules Powder Company
Hercules Research Center
Attn: Dr. Herman Skolnik, Mgr.
Technical Information Div.
Wilmington, Delaware

Hercules Powder Company
Attn: Mr. G.A. McFarren
Research Department
Wilmington 8, Delaware

Hiller Aircraft Corporation
Attn: Mr. Glenn Lattin
Chief Process Engineer
Palo Alto, California

Horizons, Inc.
Attn: Mr. M.S. Vokasovich
Ceramics Department
2905 East 79th Street
Cleveland 4, Ohio

Hughes Aircraft Company
Attn: Mr. L.E. Gates, Jr.
Advanced Structures Department
Components & Materials
Culver City, California

Hughes Aircraft Company
Attn: Mr. Glenn L. Robinson
Member of Technical Staff
Bldg. 6, Room Z 1054
Culver City, California

Hughes Tool Company
Aircraft Division
Attn: Mr. Henry G. Smith
Structures Design Staff Engr.
Bldg. 2, Station T2A
Culver City, California

Johns-Manville Fiber Glass, Inc.
Technical Center
Attn: Mr. Dominick Labino, V.P.
Director of R & D
Waterville, Ohio

Kaman Aircraft Corporation
Attn: Mr. Charles W. Ellis
Chief Test Operations Engr.
Bloomfield, Connecticut

Arthur D. Little, Inc.
Attn: Mr. D. Johnson
Research & Development
15 Acorn Park
Cambridge 40, Massachusetts

Lockheed Aircraft Corporation
Attn: Mr. H.P. Gilpin
Production Engineer
Marietta, Georgia

Lockheed Aircraft Corporation
Attn: Mr. J.F. Zechmeister, Mgr.
Test Lab Services
Dept. 70-20, Bldg. 102
Missiles & Space Division
Sunnyvale, California

Lockheed Aircraft Corporation
Attn: Mr. Elliot A. Green, Mgr.
Production Engineering
Burbank, California

Lockheed Missiles & Space Company
Attn: Mr. Roger D. Winquist
Dept. 65/10, Bldg. 152
Sunnyvale, California

The Marquardt Corporation
Attn: Mr. Joseph W. Chambers,
Supervisor
Materials & Processes Laboratories
16555 Saticoy Street
Van Nuys, California

The Martin Company
Attn: Mr. Fred D. Jewett
Structures Staff Engineer
Baltimore 3, Maryland

The Martin Company
Attn: Mr. G.C. Pfaff
Chief Structures Engineer
Orlando, Florida

AFRC (CROTLA)
L.G. Hanscom Field, Massachusetts

The Martin Company
Attn: Mr. P.M. Knox, Chief
Advanced Design
Mail Station C-109
Denver 1, Colorado

Mass. Institute of Technology
Attn: Prof. Frederick H. McGarry
Associate Professor of Materials
Dept. of Civil & Sanitary Engrg.
Cambridge 39, Massachusetts

Materials Research Corporation
Attn: Mr. V.E. Adler
Route 303
Orangeburg, New York

McDonnell Aircraft Corporation
Attn: Mr. Emil G. Szabo, Mgr.
Production Engineering
P.O. Box 516
St. Louis 3, Missouri

Minnesota Mining & Mfg. Company
Attn: Mr. R.G. Riedesel, Supervisor
Fiber Dept., New Product Div.
444 McKnight Road
Saint Paul 19, Minnesota

Mobay Chemical Company
Attn: Dr. Samuel Steingiser
Asst. Director of Research
Parkway West
Pittsburgh, Pennsylvania, 15205

National Academy of Sciences
Materials Advisory Board
Attn: Mr. John M. Wilson
2101 Constitution Avenue
Washington 25, D.C.

National Bureau of Standards
Attn: Dr. Leo A. Wall
Research Chemist, Polymer Laboratory
Washington 25, D.C.

North American Aviation, Inc.
Attn: Mr. L.P. Spalding
Section Head, Materials
Los Angeles 45, California

North American Aviation, Inc.
Attn: Mr. D.G. Worth
Engineering Department
4300 E. Fifth Avenue
Columbus 16, Ohio

North American Aviation, Inc.
Attn: Mr. B.F. Witte, Specialist
Engineering Department
Nonmetallic Materials Laboratory
International Airport
Los Angeles 45, California

Northrop Corporation
Norair Division
Attn: Mr. Royal B. Jackman, Chief
Engineering Laboratories
Hawthorne, California

Northrop-Ventura
Attn: Dr. A. Knacke
1515 Rancho Conejo Blvd.
Newbury Park, California

Owens-Corning Fiberglas Corp.
Attn: Mr. R.K. Biggers
Aero Space Division
717 - 5th Avenue
New York 22, New York

Owens-Corning Fiberglas Corporation
Attn: Mr. E.S. Cobb
Textile Prod. Development Laboratory
Ashton, Rhode Island

Owens-Corning Fiberglas Corp.
Attn: Mr. R.J. Weaver, Suite 508
900 - 17th Street, N.W.
Washington 6, D.C.

Owens-Corning Fiberglas Corp.
Attn: Mr. W.N. Edmunds
Physics Research Laboratory
Granville, Ohio

Pittsburgh Plate Glass Company
Attn: Mr. D.W. Denniston
Fiber Glass Division
Box 11422
Pittsburgh 38, Pennsylvania

Raybestos-Manhattan, Inc.
Attn: Mr. Charles P. Ellis, Jr.
Reinforced Plastics Division
1400 E. Orangethorpe Avenue
Fullerton, California

Raytheon Company, Plastic Plant
Attn: Mr. John J. Modilia
Research Division
Application Engrg. Department
Box 25
Maynard, Massachusetts

Republic Aviation Corporation
Attn: Mr. Peter L. Waters, Chief
Structures Engineer
Building 55
Farmingdale, L.I., New York

Rohr Aircraft Corporation
Attn: Mr. Hugh M. Rush
Laboratory Manager
Chula Vista, California

Ryan Aeronautical Company
Attn: Mr. J.L. Hull, Chf. Metallurgist
Materials & Processes Laboratory
San Diego 12, California

Sikorsky Aircraft Company
Attn: Mr. H.T. Jensen
Chief of Test Branch
Stratford, Connecticut

Solar Aircraft Company
Attn: Mr. Sumner Alpert, Mgr.
Aerospace & Industrial Prods.
San Diego 12, California

Southern Research Institute
Attn: Mr. W.C. Sheehan
Physical Sciences Division
2000 - 9th Avenue, South
Birmingham, Alabama

Space Age Materials Corporation
Attn: Mr. W.A. Robba
Research Department
48 - 47 58th Street
Woodside 77, New York

J.P. Stevens & Company, Inc.
Attn: Mr. R.G. Adams
Central Research Laboratory
Glass Fiber Research
141 Lanza Avenue
Garfield, New Jersey

Temco Electronics & Missiles Co.
Attn: Mr. C.L. Caudill, Sr. Gp.
Engr.

P.O. Box 6191
Dallas, Texas

H.I. Thompson Fiber Glass Co.
Attn: Mr. P. Sterry
Box 1097, Alondra Station
Gardens, California

Thomson Ramo Woodridge, Inc.
Attn: Mr. E.G. Pekarok, Mgr.
Materials Tech Services, Tabco Gp.
23555 Euclid Avenue
Cleveland 17, Ohio

United Aircraft Corporation
Attn: Mr. James F. Bacon
Materials Section
Research Laboratory
400 Main Street
East Hartford, Connecticut

United States Rubber Company
Attn: Mr. R. Lozar
Special Representative, Gov't.
Dept.

137 N. Main Street
Dayton 2, Ohio

Westinghouse Electric Corporation
Attn: Mr. Arland F. Palmer
32 N. Main Street
Dayton, Ohio

Wyandotte Chemicals Corporation
Attn: Mr. K.E. Davis
Inorganic Research
1609 Biddle Avenue
Wyandotte, Michigan

Avco Corporation
Attn: Mr. Richard J. McBride, Area Mgr.
Research & Development
Defense & Industrial Prods. Group
379 West First Street
Dayton, Ohio

Minnesota Mining & Mfg. Company
Attn: Mr. James V. Erwin
2501 Hudson Road
Saint Paul 19, Minnesota

School of Textiles
Attn: Dr. R.W. Work, Dir. of Research
North Carolina State College
Raleigh, North Carolina

Martin Company
Attn: Mr. I. Pincus, Chf.
Materials Eng. & Dev.
Baltimore, Maryland

Space-General Corp.
Attn: Mr. S.L. Lieberman, Supv.
Non-Metallics & Composites
El Monte, California

Thiokol Chemical Corporation
Attn: Mr. Hoyt Sherard
Bldg. 7611
Huntsville, Alabama

Solar Div. of International Harvester
Attn: Dr. A.G. Metcalf
2200 Pacific Highway
San Diego, California

Lundy Electronics & Systems, Inc.
Attn: Mr. James E. Henning, Asst.
Technical Dir.

Technical Center
3901 N.E. 12th Avenue
Pompano Beach, Florida

Plastics Technical Evaluation Center
Attn: SMUPA - VP-3 (A.M. Anzalone)
Picatinny Arsenal
Dover, New Jersey

P.R. Mallory & Company, Inc.
3029 E. Washington Street
Attn: Technical Library
Indianapolis, Indiana, 46206

Technical Information Systems
Attn: Ruth Christensen, Tech.
Librarian
Suttons Bay, Michigan

MAX (Dr. Lovelace)
Wright-Patterson AFB, Ohio 45433

MAN
Wright-Patterson AFB, Ohio 45433

MANL
Wright-Patterson AFB, Ohio 45433

MANE
Wright-Patterson AFB, Ohio 45433

MANP
Wright-Patterson AFB, Ohio 45433

MAAM (Mrs. Parker)
Wright-Patterson AFB, Ohio 45433

MAA
Wright-Patterson AFB, Ohio 45433

MAT
Wright-Patterson AFB, Ohio 45433

MAY
Wright-Patterson AFB, Ohio 45433

MAM
Wright-Patterson AFB, Ohio 45433

AFML (MANC)
Attn: Mr. R.C. Tomashot
Wright-Patterson AFB, Ohio ,45433

B.F. Goodrich, Research Center
Attn: Dr. William Kroeke
Brecksville, Ohio

Hercules Powder Company
Explosives and Chemical Propulsion Dept.
Attn: Mr. Donald P. Hug, Supt. Res. Dept.
P.O. Box A
Rocky Hill, New Jersey

AFML (MAA)
Technical Library
Contract AF 33(615)-1370
Wright-Patterson AFB, Ohio 45433

Ferro Corporation
Attn: Mr. R.W. Pelz
4150 East 56th Street
Cleveland 5, Ohio

Johns-Manville Corporation
Attn: Katherine L. Kinder
Documents Custodian
Research Center
Manville, New Jersey

Sandia Corporation
Attn: R.R. Sowell, 1110
Albuquerque, New Mexico

Sandia Corporation
Attn: J.M. Brierly, 8115-2
Livermore, California

Convair
Division of General Dynamics Corporation
Attn: Mr. J.F. Watson, Sr., Research Engr.
Astronautics Division
San Diego 12, California

DOCUMENT CONTROL DATA - R&D		
(Security classification of title, body of abstract and indexing annotation must be entered when the overall report is classified)		
1 ORIGINATING ACTIVITY (Corporate author) Owens-Corning Fiberglas Corporation Granville, Ohio		2a REPORT SECURITY CLASSIFICATION Unclassified
		2b GROUP
3 REPORT TITLE High Strength, High Modulus Glass Fibers		
4 DESCRIPTIVE NOTES (Type of report and inclusive dates) Technical Documentary Report March 1964-Feb. 1965		
5 AUTHOR(S) (Last name, first name, initial) McMarlin, Robert M. Tiede, Ralph L. Veazie, F. Munro		
6 REPORT DATE March 1965	7a. TOTAL NO. OF PAGES 56	7b. NO. OF IEP'S 5
8a. CONTRACT OR GRANT NO. AF 33(615)-1370	9a. ORIGINATOR'S REPORT NUMBER(S) None	
b. PROJECT NO. 7320		
c. 7340	9b. OTHER REPORT NO(S) (Any other numbers that may be assigned this report) AFML-TR-65- Part I	
d.		
10. AVAILABILITY/LIMITATION NOTICES Qualified requesters may obtain copies of this report from the Defense Documentation Center (DDC), (formerly ASTIA), Cameron Station, Bldg. 5, 5010 Duke Street, Alexandria 4, Virginia. Dissemination outside the Department of Defense or to recipients other than Government defense contractors is prohibited. DDC release to CFSII not authorized.		
11. SUPPLEMENTARY NOTES None		12. SPONSORING MILITARY ACTIVITY AFML-RTD-AFSC-USAF
13. ABSTRACT The results of the present investigation have shown that glass fibers drawn from compositions located in the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-ZnO}$ system possess tensile strengths in excess of 800,000 psi. Glass fibers drawn from compositions in the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-MgO}$ system have given similar results, particularly at fiber diameters in the 0.00015 to 0.00020 inch range. The modulus of elasticity of glass fibers obtained from compositions located in the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-MgO}$ system can be increased to 16.0×10^6 psi. This is accomplished by lowering the fiber diameter to the 0.00015 to 0.00020 inch range. In the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-Na}_2\text{O}$ system, the fiber forming temperature, the tensile strength and the modulus of elasticity change with composition in a manner which gives a break in the composition vs. property curves at the point where the Al to Na ratio is 1. This is related to the change in the coordination number of the aluminum at this ratio which has been proposed in the literature. The properties measured for glasses in the $\text{SiO}_2\text{-TiO}_2\text{-K}_2\text{O}$ system do not appear to lead to as clear a concept of glass structure. Properties very gradually with composition, with no evidence of a sharp break at any point. There is some indication that the rate of change of properties is dominated more by the silica content than is the case in the $\text{SiO}_2\text{-Al}_2\text{O}_3\text{-Na}_2\text{O}$ system.		

KEY WORDS

Fibers, glass, composition, strength, modulus

LINK A

LINK B

LINK C

ROLE

WT

ROLE

WT

ROLE

WT

INSTRUCTIONS

1. **ORIGINATING ACTIVITY:** Enter the name and address of the contractor, subcontractor, grantee, Department of Defense activity or other organization (*corporate author*) issuing the report.

2a. **REPORT SECURITY CLASSIFICATION:** Enter the overall security classification of the report. Indicate whether "Restricted Data" is included. Marking is to be in accordance with appropriate security regulations.

2b. **GROUP:** Automatic downgrading is specified in DoD Directive 5200.10 and Armed Forces Industrial Manual. Enter the group number. Also, when applicable, show that optional markings have been used for Group 3 and Group 4 as authorized.

3. **REPORT TITLE:** Enter the complete report title in all capital letters. Titles in all cases should be unclassified. If a meaningful title cannot be selected without classification, show title classification in all capitals in parentheses immediately following the title.

4. **DESCRIPTIVE NOTES:** If appropriate, enter the type of report, e.g., interim, progress, summary, annual, or final. Give the inclusive dates when a specific reporting period is covered.

5. **AUTHOR(S):** Enter the name(s) of author(s) as shown on, or in the report. Enter last name, first name, middle initial. If military, show rank and branch of service. The name of the principal author is an absolute minimum requirement.

6. **REPORT DATE:** Enter the date of the report as day, month, year; or month, year. If more than one date appears on the report, use date of publication.

7a. **TOTAL NUMBER OF PAGES:** The total page count should follow normal pagination procedures, i.e., enter the number of pages containing information.

7b. **NUMBER OF REFERENCES:** Enter the total number of references cited in the report.

8a. **CONTRACT OR GRANT NUMBER:** If appropriate, enter the applicable number of the contract or grant under which the report was written.

8b, 8c, & 8d. **PROJECT NUMBER:** Enter the appropriate military department identification such as project number, subproject number, system numbers, task number, etc.

9a. **ORIGINATOR'S REPORT NUMBER(S):** Enter the official report number by which the document will be identified and controlled by the originating activity. This number must be unique to this report.

9b. **OTHER REPORT NUMBER(S):** If the report has been assigned any other report numbers (either by the originator or by the sponsor), also enter this number(s).

10. **AVAILABILITY/LIMITATION NOTICES:** Enter any limitations on further dissemination of the report, other than those

imposed by security classification, using standard statements such as:

- (1) "Qualified requesters may obtain copies of this report from DDC."
- (2) "Foreign announcement and dissemination of this report by DDC is not authorized."
- (3) "U. S. Government agencies may obtain copies of this report directly from DDC. Other qualified DDC users shall request through _____."
- (4) "U. S. military agencies may obtain copies of this report directly from DDC. Other qualified users shall request through _____."
- (5) "All distribution of this report is controlled. Qualified DDC users shall request through _____."

If the report has been furnished to the Office of Technical Services, Department of Commerce, for sale to the public, indicate this fact and enter the price, if known.

11. **SUPPLEMENTARY NOTES:** Use for additional explanatory notes.

12. **SPONSORING MILITARY ACTIVITY:** Enter the name of the departmental project office or laboratory sponsoring (paying for) the research and development. Include address.

13. **ABSTRACT:** Enter an abstract giving a brief and factual summary of the document indicative of the report, even though it may also appear elsewhere in the body of the technical report. If additional space is required, a continuation sheet shall be attached.

It is highly desirable that the abstract of classified reports be unclassified. Each paragraph of the abstract shall end with an indication of the military security classification of the information in the paragraph, represented as (TS) (S) (C) or (U).

There is no limitation on the length of the abstract. However, the suggested length is from 150 to 225 words.

14. **KEY WORDS:** Key words are technically meaningful terms or short phrases that characterize a report and may be used as index entries for cataloging the report. Key words must be selected so that no security classification is required. Identifiers, such as equipment model designation, trade name, military project code name, geographic location, may be used as key words but will be followed by an indication of technical content. The assignment of links, rules, and weights is optional.